

## Periodic Trends Guided Note Sheet

Name \_\_\_\_\_

Period \_\_\_\_\_

Date \_\_\_\_\_

Fill in the blanks. Some blanks have word choices written after them to help you. Refer to the drawings and questions from the first page to help you with the fill in the other pages.

### I. Atomic Radius in a Group

- A. As you move down a group, the nucleus becomes \_\_\_\_\_ (stronger, weaker) because more \_\_\_\_\_ (electrons, protons) are being added to it. This stronger nucleus can pull the electrons in more tightly, making the atoms **smaller**; **smaller**.
- B. \_\_\_\_\_ (Protons, Electrons) are also added to the atom but are added to \_\_\_\_\_ (larger, smaller, the same) energy levels resulting in **larger** atoms.
- C. The increased number of inner electrons shield (or block) the \_\_\_\_\_ (positive, negative) pull of the nucleus.
- D. So, electron shielding \_\_\_\_\_ (increases, decreases) causing the electrons on the outermost energy levels to be even farther from the nucleus than expected.
- E. This causes the atoms at the bottom of a Group to be even \_\_\_\_\_ (larger, smaller) than expected.
- F. Therefore, atomic radius \_\_\_\_\_ (increases, decreases) as you move down a group.

### II. Atomic Radius in a Period

- A. As you move right across a Period, the nucleus becomes \_\_\_\_\_ (stronger, weaker) because more \_\_\_\_\_ (electrons, protons) are being added to it.
- B. \_\_\_\_\_ (Protons, Electrons) are also added but they are added to \_\_\_\_\_ (the same, a larger, a smaller) energy level.
- C. This causes the amount of electron shielding from electrons on inner energy levels to remain \_\_\_\_\_ (constant, varied).
- D. Therefore, electrons are pulled in \_\_\_\_\_ (more, less) tightly because they feel the pull from nucleus more strongly.
- E. This causes the atoms on the right of the Periodic Table to be \_\_\_\_\_ (larger, smaller) than those on the left.
- F. So, atomic radius \_\_\_\_\_ (increases, decreases) as you move across a Period from left to right.

### III. Sizes of an Ion Versus Sizes of an Atom (Digital Textbook—Section 11.2)

- A. Metal cations are \_\_\_\_\_ (smaller, larger) than the neutral atoms.
- B. because \_\_\_\_\_
- C. Nonmetal anions are \_\_\_\_\_ (smaller, larger) than the neutral atoms.
- D. because \_\_\_\_\_

### IV. Ionization Energy in a Group

- A. Ionization energy is the energy required to \_\_\_\_\_ the outermost electron from an atom.
- B. As you move down a Group, the nucleus becomes \_\_\_\_\_ (stronger, weaker) because there are \_\_\_\_\_ (more, fewer) protons in it.
- C. **Constant**, as you move down a group, there are \_\_\_\_\_ (fewer, more) energy levels making the atom \_\_\_\_\_ (larger, smaller).
- D. In atoms at the top of a group, the outermost electrons are closer to the nucleus so the electrons can feel the pull of nucleus \_\_\_\_\_ (more, less) strongly.
- E. The outermost electrons also feel the pull of the nucleus less strongly because there is \_\_\_\_\_ (less, more) electron shielding from the \_\_\_\_\_ (decreased, increased) number of inner electrons.
- F. Therefore, \_\_\_\_\_ (more, less) energy is required to remove the outermost electron from a larger atom.

# Periodic Trends Review Notes Sheet Answers

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## **Periodic Trends Review Notes Sheet Answers:**

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## **Table of Contents Periodic Trends Review Notes Sheet Answers**

1. Understanding the eBook Periodic Trends Review Notes Sheet Answers
  - The Rise of Digital Reading Periodic Trends Review Notes Sheet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Review Notes Sheet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periodic Trends Review Notes Sheet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Review Notes Sheet Answers
  - Personalized Recommendations
  - Periodic Trends Review Notes Sheet Answers User Reviews and Ratings
  - Periodic Trends Review Notes Sheet Answers and Bestseller Lists
5. Accessing Periodic Trends Review Notes Sheet Answers Free and Paid eBooks
  - Periodic Trends Review Notes Sheet Answers Public Domain eBooks
  - Periodic Trends Review Notes Sheet Answers eBook Subscription Services
  - Periodic Trends Review Notes Sheet Answers Budget-Friendly Options

6. Navigating Periodic Trends Review Notes Sheet Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Periodic Trends Review Notes Sheet Answers Compatibility with Devices
  - Periodic Trends Review Notes Sheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Periodic Trends Review Notes Sheet Answers
  - Highlighting and Note-Taking Periodic Trends Review Notes Sheet Answers
  - Interactive Elements Periodic Trends Review Notes Sheet Answers
8. Staying Engaged with Periodic Trends Review Notes Sheet Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Periodic Trends Review Notes Sheet Answers
9. Balancing eBooks and Physical Books Periodic Trends Review Notes Sheet Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Periodic Trends Review Notes Sheet Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Periodic Trends Review Notes Sheet Answers
  - Setting Reading Goals Periodic Trends Review Notes Sheet Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Trends Review Notes Sheet Answers
  - Fact-Checking eBook Content of Periodic Trends Review Notes Sheet Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements



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