

Periodic Trends Guided Note Sheet

Name _____

Period _____

Date _____

Fill in the blanks. Some blanks have word choices written after them to help you. Refer to the drawings and questions from the first page to help you with the fill in the other pages.

I. Atomic Radius in a Group

- As you move down a group, the nucleus becomes _____ (stronger, weaker) because more _____ (electrons, protons) are being added to it. This stronger nucleus can pull the electrons in more tightly, making the atoms **smaller**; **smaller**.
- _____ (Protons, Electrons) are also added to the atom but are added to _____ (larger, smaller, the same) energy levels resulting in **larger** atoms.
- The increased number of inner electrons shield (or block) the _____ (positive, negative) pull of the nucleus.
- So, electron shielding _____ (increases, decreases) causing the electrons on the outermost energy levels to be even farther from the nucleus than expected.
- This causes the atoms at the bottom of a Group to be even _____ (larger, smaller) than expected.
- Therefore, atomic radius _____ (increases, decreases) as you move down a group.

II. Atomic Radius in a Period

- As you move right across a Period, the nucleus becomes _____ (stronger, weaker) because more _____ (electrons, protons) are being added to it.
- _____ (Protons, Electrons) are also added but they are added to _____ (the same, a larger, a smaller) energy level.
- This causes the amount of electron shielding from electrons on inner energy levels to remain _____ (constant, varied).
- Therefore, electrons are pulled in _____ (more, less) tightly because they feel the pull from nucleus more strongly.
- This causes the atoms on the right of the Periodic Table to be _____ (larger, smaller) than those on the left.
- So, atomic radius _____ (increases, decreases) as you move across a Period from left to right.

III. Sizes of an Ion Versus Sizes of an Atom (Digital Textbook—Section 11.2)

- Metal cations are _____ (smaller, larger) than the neutral atoms.
- because _____
- Nonmetal anions are _____ (smaller, larger) than the neutral atoms.
- because _____

IV. Ionization Energy in a Group

- Ionization energy is the energy required to _____ the outermost electron from an atom.
- As you move down a Group, the nucleus becomes _____ (stronger, weaker) because there are _____ (more, fewer) protons in it.
- Constant**, as you move down a group, there are _____ (fewer, more) energy levels making the atoms _____ (larger, smaller).
- In atoms at the top of a group, the outermost electrons are closer to the nucleus so the electrons can feel the pull of nucleus _____ (more, less) strongly.
- The outermost electrons also feel the pull of the nucleus less strongly because there is _____ (less, more) electron shielding from the _____ (decreased, increased) number of inner electrons.
- Therefore, _____ (more, less) energy is required to remove the outermost electron from a larger atom.

Periodic Trends Review Notes Sheet Answers

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