

MYP4 Chemistry

Periodic Trends

Use the periodic table, charts, and your knowledge of periodic trends to answer the following questions.

Atomic Radius

1. Which atom in each pair has the larger atomic radius? Circle your answer.

a) Li or **K** b) **Ca** or Ni c) **Ga** or B d) O or **C** e) Cl or **Br**
f) Be or **Ba** g) **Si** or S h) Fe or **Au**

2. What is the periodic trend for atomic size from top to bottom in a group? From left to right in a period?

Atomic size (atomic radius) increases from top to bottom in a group and decreases from left to right in a period.

3. Why do atoms get smaller as you move left to right in a period?

Atoms get smaller as you move left to right in a period due to the increase in effective nuclear charge (while the number of occupied energy levels remain constant) leading to an increase in the force of attraction, which pulls the outer electrons closer to the nucleus and decreases the atomic radius.

4. Why do atoms get larger as you move down a group?

Atoms get larger as you move down a group due to the increase in occupied energy levels.

5. Arrange the following in order of increasing atomic radius atomic radius.

Mg, Na, P, Si and Ar? They are in the same period: Ar, P, Si, Mg, Na

Ionic Radius

6. Define an ion.

a charged atom or group of atoms

7. What is the difference between a cation and an anion?

Cations are positively charged ions. They are formed when an atom loses electrons.
Anions are negatively charged ions. They are formed when an atom gains electrons.

8. Explain the relationship between the relative size of an ion to its neutral atom and the charge on the ions.

Cations are smaller than their neutral atom because they have less occupied energy levels.
Anions are larger than their neutral atom because of the increase in electron-electron repulsion.

9. Which particle has the larger radius in each atom/ion pair? Circle your answer.

a) **Na** Na⁺ b) S, **S²⁻** c) I, I⁻ d) **Al** Al³⁺

Periodic Trends Practice Answer Key

**Nelson, C.R., Hallett, J.G., Romero
Montoya, A.E., Andrade, A., Besacier,
C., Boerger, V., Bouazza, K., Chazdon,
R., Cohen-Shacham, E., Danano,
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M., Guariguata, M.R., Gutierrez,
V., Hancock, B., Innecken, P., Katz,
S.M., McCormick, R., Moraes,
L.F.D., Murcia, C., Nagabhatla,
N., Pouaty Nzembialela, D., Rosado-
May, F.J., Shaw, K., Swiderska,
K., Vasseur, L., Venkataraman,**

**R., Walder, B., Wang, Z., Weidlich,
E.W.A.**

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Problems and Problem Solving in Chemistry Education Georgios Tsaparlis, 2021-05-17

Problem solving is central to the teaching and learning of chemistry at secondary, tertiary and post tertiary levels of education, opening to students and professional chemists alike a whole new world for analysing data, looking for patterns and making deductions. As an important higher order thinking skill, problem solving also constitutes a major research field in science education. Relevant education research is an ongoing process with recent developments occurring not only in the area of quantitative computational problems but also in qualitative problem solving. The following situations are considered: some general others with a focus on specific areas of chemistry: quantitative problems, qualitative reasoning, metacognition and resource activation, deconstructing the problem solving process, an overview of the working memory hypothesis, reasoning with the electron pushing formalism, scaffolding organic synthesis skills, spectroscopy for structural characterization in organic chemistry, enzyme kinetics, problem solving in the academic chemistry laboratory, chemistry problem solving in context, team based active learning, technology for molecular representations, IR spectra simulation and computational quantum chemistry tools. The book concludes with methodological

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