

Concepts in Chemistry

The Periodic Law

SECTION 2

Select the correct answer for the following questions for the space provided.

1. _____ for the modern periodic table, elements are ordered
 - a) according to increasing atomic mass.
 - b) according to Mendeleev's original design.
 - c) according to increasing atomic number.
 - d) based on how they were discovered.
2. _____ Mendeleev stated that certain elements have similar chemical properties
 - a) elements appeared in regular intervals and elements discovered later completed intervals of his spacing.
 - b) elements.
 - c) atomic numbers.
 - d) mass ratios.
3. _____ The modern periodic law states that
 - a) all new elements with the same properties are listed in the same group.
 - b) the physical and chemical properties of elements are functions of atomic number.
 - c) elements exhibit properties of both periods and groups.
 - d) the chemical properties of elements can be grouped according to periodicity but physical properties cannot.
4. _____ The discovery of the noble gases changed Mendeleev's periodic table
 - a) adding a row.
 - b) adding a column.
 - c) group.
 - d) none.
5. _____ The noble gases that property of the noble gases is that they are
 - a) unstable.
 - b) inert.
 - c) unstable.
 - d) not very reactive.
6. _____ Likewise, the first column in Group 1, has an atomic number of 3. The second element in this group has an atomic number of
 - a) 4.
 - b) 11.
 - c) 19.
 - d) 20.
7. How groups in Mendeleev's periodic table of 18 and the atomic number of 55
 - a) 2. How many elements are in this group?
 - b) 2. How many elements are in this group?
 - c) 2. How many elements are in this group?
 - d) 2. How many elements are in this group?

Periodic Law Chapter 5 Test Answers

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Periodic Law Chapter 5 Test Answers:

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