CHAPTER 5 REVIEW

The Periodic Law

SECTION 1

SHORT ANSWER   Answer the following questions in the space provided.

1. ___c___ In the modern periodic table, elements are ordered

  (a) according to decreasing atomic mass.
  (b) according to Mendeleev's original design.
  (c) according to increasing atomic number.
  (d) based on when they were discovered.

2. ___d___ Mendeleev noticed that certain similarities in the chemical properties of elements appeared at regular intervals when the elements were arranged in order of increasing

  (a) density.   **(c)** atomic number.
  (b) reactivity.   **(d)** atomic mass.

3. ___b___ The modern periodic law states that

  (a) no two electrons with the same spin can be found in the same place in an atom.
  (b) the physical and chemical properties of an element are functions of its atomic number.
  (c) electrons exhibit properties of both particles and waves.
  (d) the chemical properties of elements can be grouped according to periodicity, but physical properties cannot.

4. ___c___ The discovery of the noble gases changed Mendeleev's periodic table by adding a new

  (a) period.   **(c)** group.
  (b) series.   **(d)** level.

5. ___d___ The most distinctive property of the noble gases is that they are

  (a) metallic.   **(c)** metalloid.
  (b) radioactive.   **(d)** largely unreactive.

6. ___c___ Lithium, the first element in Group 1, has an atomic number of 3. The second element in this group has an atomic number of

  (a) 4.   **(c)** 11.
  (b) 10.   **(d)** 18.

7. An isotope of fluorine has a mass number of 19 and an atomic number of 9.

  ___9___ **a.** How many protons are in this atom?

  ___10___ **b.** How many neutrons are in this atom?

  ___$^{19}_{9}F$___ **c.** What is the nuclear symbol of this fluorine atom, including its mass number and atomic number?

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