Name: Key

Date:

Honors Chemistry

Percent Yield Practice

 Calculate the theoretical yield of calcium oxide if 24.8 g of calcium carbonate decomposes. Calculate the percent yield if you got 13.1 g in the laboratory.

24.19 (aco = 13.1 100) = 94.2490 => 94.296

2. What is the percent yield if you experimentally produce 3.65 g of copper when 1.87 g of aluminum reacts with 9.65 g copper (II) sulfate?

2 Al + 3 Cuso, - 3 Cu + Alz(sou),

1.87 Al . 1 met Al . 3 met Cuso, 159.629 (uson = 16.69 Cuso,

26.98 Al . 2 met Al . 1 met Cuso, 1 met C

3. Fyou start with 5.0 g of C₃H₈ and react it with oxygen, what is the theoretical yield of water? If you have a percent yield of 75.0%, What is the actual yield of water?

2 _C,H₈ + 5 O₂ - 3 CO₂ + 4 H₂O

5.09 GHs. | mot C,H₅ | 4 mot H₂O | 18.02 gH₂O = 8.22 gH₂O |

PY = 1V 100

PY - 1V 100

PY - TY = AY = 75 L 8.22 = 6.12 g H₂O

4. What is the theoretical yield of sodium oxide if you start with 20 g of calcium oxide?

ZNaCI + _ CaO - _ CaCI2 + Na2O

ZOg CaO . Inst CaO . Inst tac 61.98 g Maz 0 = 200 g Maz 0

Percent Yield Answer Key

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