

# Percent Yield Practice

3 pages + 3-page Answer KEY + Periodic Table

Chemistry

## PERCENT YIELD

Show all work and circle or highlight your final answer.

$$\text{percent yield (\%)} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

- 1 16.8 grams of sodium reacts with excess chlorine gas.



The theoretical yield of sodium chloride is 25.5 grams. The actual yield was 23.5 grams. Calculate the percent yield of sodium chloride.

- 2 4.8 grams of carbon reacts with excess oxygen gas.



The theoretical yield of carbon dioxide is 25.5 g. Calculate the percent yield of carbon dioxide.

- 3 1.6 grams of zinc reacts with excess sulfur.



The theoretical yield of zinc sulfide is 11.5 grams. Calculate the percent yield of zinc sulfide.

- 4 1.0 grams of phosphorus reacts with excess chlorine.



The theoretical yield of phosphorus trichloride is 4.0 grams. Calculate the percent yield of phosphorus trichloride.

- 5 Hydrogen gas ( $\text{H}_2$ ) reacts with oxygen gas ( $\text{O}_2$ ) to produce water ( $\text{H}_2\text{O}$ ). 5.8 grams of hydrogen reacts with excess oxygen gas.



What is the theoretical yield?

- 6 The actual yield was 16.8 grams. Calculate the percent yield.

- 7 Magnesium ( $\text{Mg}$ ) reacts with hydrochloric acid ( $\text{HCl}$ ) to produce magnesium chloride ( $\text{MgCl}_2$ ) and hydrogen gas ( $\text{H}_2$ ). 1.0 gram of magnesium reacts with excess hydrochloric acid.



What is the theoretical yield?

- 8 The actual yield was 40.2 grams. Calculate the percent yield.

- 9 Calcium carbonate ( $\text{CaCO}_3$ ) reacts with hydrochloric acid ( $\text{HCl}$ ) to produce calcium chloride ( $\text{CaCl}_2$ ), carbon dioxide ( $\text{CO}_2$ ), and water ( $\text{H}_2\text{O}$ ). 15.8 grams of calcium carbonate reacts with excess hydrochloric acid.



What is the theoretical yield of calcium chloride?

$$\begin{array}{r} 15.8 \text{ g CaCO}_3 \times \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3} \times \frac{1 \text{ mol CaCl}_2}{1 \text{ mol CaCO}_3} \times \frac{146.97 \text{ g CaCl}_2}{1 \text{ mol CaCl}_2} \\ = 0.158 \text{ mol CaCO}_3 \times 1 \text{ mol CaCl}_2 \\ = 0.158 \text{ mol CaCl}_2 \end{array}$$

$$\begin{array}{r} 0.158 \text{ mol CaCl}_2 \times \frac{1 \text{ mol CaCl}_2}{1 \text{ mol CaCl}_2} \times \frac{146.97 \text{ g CaCl}_2}{1 \text{ mol CaCl}_2} \\ = 0.158 \text{ mol CaCl}_2 \times 146.97 \text{ g CaCl}_2 \\ = 23.22 \text{ g CaCl}_2 \end{array}$$

$$\begin{array}{r} \text{Molar Mass CaCl}_2 \\ \text{Ca} = 40.08 \\ \text{Cl} = 35.45 \\ 2 \times 35.45 = 70.90 \\ 40.08 + 70.90 = 110.98 \text{ g} \end{array}$$

$$\begin{array}{r} \text{Molar Mass CaCl}_2 \\ \text{Ca} = 40.08 \\ \text{Cl} = 35.45 \times 2 = 70.90 \\ 40.08 + 70.90 = 110.98 \text{ g} \end{array}$$

- 10 The actual yield was 22.5 grams. Calculate the percent yield of calcium chloride.

$$\text{percent yield (\%)} = \frac{22.5 \text{ g}}{23.22 \text{ g}} \times 100 = 9.70\% \times 100 = 97.0\% \text{ CaCl}_2$$

- 11 Potassium ( $\text{K}$ ) reacts with iodine ( $\text{I}_2$ ) to produce potassium iodide ( $\text{KI}$ ). 1.8 grams of potassium reacts with excess iodine.



What is the theoretical yield?

$$\begin{array}{r} 1.8 \text{ g K} \times \frac{1 \text{ mol K}}{39.10 \text{ g K}} \times \frac{2 \text{ mol KI}}{2 \text{ mol K}} \times \frac{166.10 \text{ g KI}}{1 \text{ mol KI}} \\ = 0.046 \text{ mol K} \times 2 \text{ mol KI} = 0.092 \text{ mol KI} \end{array}$$

$$\begin{array}{r} \text{Molar Mass KI} \\ \text{K} = 39.10 \\ \text{I} = 126.90 \\ 39.10 + 126.90 = 166.00 \text{ g} \end{array}$$

$$\begin{array}{r} \text{Molar Mass KI} \\ \text{K} = 39.10 \\ \text{I} = 126.90 \\ 39.10 + 126.90 = 166.00 \text{ g} \end{array}$$

ANSWER KEY INCLUDED

%  
YIELD

# Percent Yield Example Answer Key

**KJ Lindholm-Leary**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

**Percent Yield Example Answer Key:**

## Unveiling the Magic of Words: A Review of "**Percent Yield Example Answer Key**"

In some sort of defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their ability to kindle emotions, provoke contemplation, and ignite transformative change is actually awe-inspiring. Enter the realm of "**Percent Yield Example Answer Key**," a mesmerizing literary masterpiece penned with a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve in to the book is central themes, examine its distinctive writing style, and assess its profound affect the souls of its readers.

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web 1 john 6 1 4 a crowd gathers to jesus near the sea of galilee after these things jesus went over the sea of galilee which is the sea of tiberias then a great multitude

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