

Answers to Problems
Introduction to Spectroscopy, 3rd edition

Pavia, Lampman and Kriz

CHAPTER 1

- (a) 90.50% carbon; 9.50% hydrogen (b) C_4H_8
- 32.0% carbon; 5.4% hydrogen; 62.8% chlorine; $C_3H_5Cl_2$
- $C_2H_3NO_2$
- 180.2 = molecular mass. Molecular formula is $C_9H_8O_4$.
- Equivalent weight = 52.3
- (a) 6 (b) 1 (c) 3 (d) 6 (e) 12
- The index of hydrogen deficiency = 1. There cannot be a triple bond, since the presence of a triple bond would require an index of hydrogen deficiency of at least 2.
- (a) 59.96% carbon; 5.75% hydrogen; 34.29% oxygen (b) $C_7H_8O_3$
(c) $C_{21}H_{24}O_8$ (d) A maximum of two aromatic (benzenoid) rings
- (a) $C_8H_8O_2$ (b) $C_8H_{12}N_2$ (c) $C_7H_8N_2O$ (d) $C_5H_{12}O_4$
- Molecular formula = $C_8H_{10}N_4O_2$; index of hydrogen deficiency = 6
- Molecular formula = $C_{21}H_{30}O_2$; index of hydrogen deficiency = 7
- For the hydrolysis product: molecular formula = $C_6H_{12}O_6$; index of hydrogen deficiency = 1

For the original carbohydrate: molecular formula = $C_{12}H_{22}O_{11}$; index of hydrogen deficiency = 2

CHAPTER 2

- (a) Propargyl chloride (3-chloropropyne)
(b) *p*-Cymene (4-isopropyltoluene)

Pavia Spectroscopy Edition Answer

N Colangelo



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