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Practice Set 2

Paper Reference **PRACL2/N02**

Mathematics
Level 2

You must have:

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**Katie Braid, Jane Ellingham, Christopher
Lindle, Hayley Thompson**

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using the equation from problem 2 above answer the following

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$$\begin{aligned} &1 \text{ n}_2 + 3 \text{ f}_2 \rightarrow 2 \text{ nf}_3 \\ &2 \text{ c}_6\text{h}_{10} + 17 \text{ o}_2 \rightarrow 12 \text{ co}_2 + 10 \text{ h}_2\text{o} \\ &3 \text{ hbr} + 1 \text{ khco}_3 + 1 \text{ h}_2\text{o} + 1 \text{ kbr} + 1 \text{ co}_2 + 2 \text{ ga}_2\text{so}_3 + 3 \text{ na}_2\text{so}_3 + 1 \text{ ga}_2\text{so}_3 + 3 \text{ so}_3 + 6 \text{ nabr} + 3 \text{ sno} + 2 \text{ nf}_3 + 3 \text{ snf}_2 + 1 \text{ n}_2\text{o}_3 \end{aligned}$$
using the following equation
$$2 \text{ naoh} + \text{h}_2\text{so}_4 \rightarrow 2 \text{ h}_2\text{o} + \text{na}_2\text{so}_4$$

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$$\text{ce} + \text{zn} + \text{cucl}_2 \rightarrow \text{zncl}_2 + \text{cu} + \text{zn} + \text{cucl}_x$$
$$2 \text{ zncl}_x + 2 \text{ cu} \rightarrow \text{how many moles of ce} + 2 \text{ zncl}_2 + 2 \text{ zn}$$
will be produced from 23.0 g of ce and 23.0 g of zn assuming ce and cucl_2 is available in excess

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$$\begin{aligned} &1 \text{ n}_2 + 3 \text{ f}_2 \rightarrow 2 \text{ nf}_3 \\ &2 \text{ c}_6\text{h}_{10} + 17 \text{ o}_2 \rightarrow 12 \text{ co}_2 + 10 \text{ h}_2\text{o} \\ &3 \text{ hbr} + 1 \text{ khco}_3 + 1 \text{ h}_2\text{o} + 1 \text{ kbr} + 1 \text{ co}_2 + 2 \text{ ga}_2\text{so}_3 + 3 \text{ na}_2\text{so}_3 + 1 \text{ ga}_2\text{so}_3 + 3 \text{ so}_3 + 6 \text{ nabr} + 3 \text{ sno} + 2 \text{ nf}_3 + 3 \text{ snf}_2 + 1 \text{ n}_2\text{o}_3 \end{aligned}$$
br using the following equation
$$2 \text{ naoh} + \text{h}_2\text{so}_4 \rightarrow 2 \text{ h}_2\text{o} + \text{na}_2\text{so}_4$$
br

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