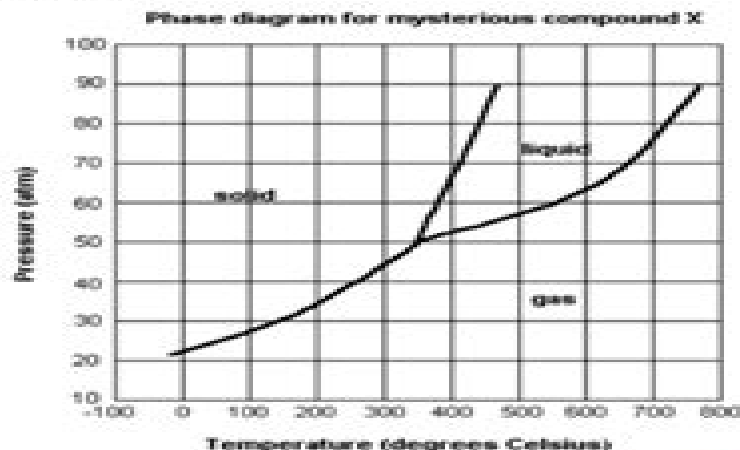


Phase Diagram Worksheet

For each of the questions on this worksheet, refer to the phase diagram for mysterious compound X.



- 1) What is the critical temperature of compound X? **$\sim 470^{\circ}\text{C}$**
- 2) If you were to have a bottle containing compound X in your closet, what phase would it most likely be in?
Extrapolating from this diagram, it's most likely a gas.
- 3) At what temperature and pressure will all three phases coexist?
 350°C , $\sim 51\text{ atm}$
- 4) If I have a bottle of compound X at a pressure of 45 atm and temperature of 100°C , what will happen if I raise the temperature to 400°C ?
It will sublime
- 5) Why can't compound X be boiled at a temperature of 200°C ?
It does not form a liquid at this temperature. It only exists as a liquid at temperatures above 350°C .
- 6) If I wanted to, could I drink compound X?
No. At the temperatures and pressures that it forms a liquid, you'd probably die.

Phase Diagram Practice Answer Key

**Theodore L. Brown, H. Eugene LeMay
Jr., Bruce E. Bursten, Catherine
Murphy, Patrick Woodward, Steven
Langford, Dalius Sagatys, Adrian
George**

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