- Give a recipe for preparing a sodium carbonate/sodium bicarbonate buffer solution having a pH of 9.5.For carbonic acid, H<sub>2</sub>CO<sub>3</sub>, pKa1 = 6.35 and pKa2 = 10.33
- 8. A saturated solution of silver chromate ( $Ag_2CrO_4$ , a sparingly soluble solid, formula mass 331.0 g/mol), was found to have a chromate ion concentration of 3.2 x  $10^{-4}$  M at 25 °C. What is (a) the  $K_{sp}$ ; and (b) solubility of the salt in mg/mL? How could you dissolve a sample of the salt?
- 9. Calculate the solubility in mg/mL of PbCl<sub>2</sub> (formula mass 276.2 g/mol) in 0.25 M sodium chloride solution at 25 °C.  $K_{sp}$  for lead chloride is 1.8 x 10-10 at 25 °C.
- 10. Write an equation to show the formation of the complex ion  $[Ag(NH_3)_2^+]$  from silver ion and aqueous ammonia then calculate the concentration of silver ions in solution if exactly 0.100 mol of silver nitrate is added to exactly 500 mL of 0.800 M ammonia solution.  $K_f$  for  $[Ag(NH_3)_2]^+$  is 1.7 x 107.

# Recipe For Sodium Carbonate Sodium Bicarbonate Buffer

Naeem Sarwar, Atique-ur-Rehman, Shakeel Ahmad, Mirza Hasanuzzaman

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