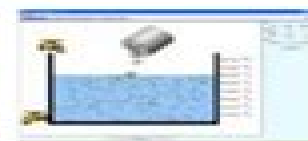


Solubility of Salts

Objective

When we cook pasta, we frequently dissolve table salt in water. Is there a limit to how much salt we can get into the water? When ionic-bonded compounds (like table salt, NaCl) are dissolved in polar solvents (like water), the **polar** water molecules act to pull the ions in the compound apart. When this happens (called dissociation), ions are released into the solution. Is there a limit to the number of ions that can exist in solution?



Salts & Solubility

Directions

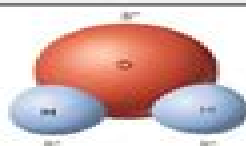
PART 1: Observation of Salt and Water Interaction

- Click on the "Table Salt" tab. [Table Salt](#)
- Begin by shaking salt shaker a few times. Make at least 2 observations of salt as it enters the water.

<http://phet.colorado.edu/>

Play with the Sims → Chemistry → Salts & Solubility [Run Now!](#)
If a yellow bar drops down in your browser, click on it and select "Allow Blocked Content"

Given the Structure for Water and IONic Salt Below... Why is Water Able To Dissolve IONic Salts?



It is able to dissolve salts because the sodium and chloride ions break apart and are able to dissolve when there is enough water.



PART 2: Solubility Limit?

- Reset the application and investigate if there is a point at which solute can no longer dissolve (called *Saturated Point*). Record the amount of ions dissolved and the Original volume of solvent used.
- Remove $\frac{1}{2}$ of the solvent from the original amount. Investigate the *saturation point* at this amount of solvent. Record the amount of ions dissolved and the volume of solvent used.
- Add $\frac{1}{2}$ more solvent to the original amount. Investigate the *saturation point* at this amount of solvent used. Record the amount of ions dissolved and the volume of solvent used.

How Does the Amount of Solvent Affect Solubility?

When there isn't enough solvent then the solubility increases because the ions are not be able to dissolve.

Maximum amount of ions dissolved	Volume at Saturation Point
Na ⁺ 181	5.0 x 10 ⁻²³ L
Cl ⁻ 181	
Na ⁺ 92	2.5 x 10 ⁻²³ L
Cl ⁻ 92	
Na ⁺ 104	5.0 x 10 ⁻²³ L
Cl ⁻ 104	

PART 3: Can Solubility Change With Other Solutes

- Click on the "Slightly Soluble Salts" tab. [Slightly Soluble Salts](#)
- Start with any solute and investigate the saturation point using the given amount of solvent. Write the formula for the solute and record the amount of ions dissolved and the volume of solvent used.
- RESET the application and repeat step 2 for two more different solutes.

Solubility of Different Solutes

Name	Strontium Phosphate			Name	Silver Bromide			Name	Silver Arsenate		
Formula	Sr3(PO4) 2			Formula	AgBr			Formula	Ag3AsO4		
Dissolved ions	Sr 45	1.00	L	Dissolved ions	Ag 44	1.00	L	Dissolved ions	Ag 252	1.00	L
	PO 30				Br 44				AsO 82		

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Albert A Gayle



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