

## I. Fill in the blanks

Mendeleev arranged the periodic table in order of increasing atomic mass. He was able to use his periodic table to predict properties of the missing elements.

Moseley discovered that each element has a unique atomic number and arranged the elements in order of increasing atomic number.

The modern periodic law states that the properties of the elements repeat when arranged by increasing atomic number.

A column on the periodic table is called a family. Elements in the same column of the periodic table have similar electron distributions. Elements in the same column of the periodic table have the same number of valence electrons.

A row on the periodic table is called a period. Elements in the same row of the periodic table have the same number of energy levels.

Elements in the Noble Gas family are considered stable because they have full (filled) outer energy levels.

Elements that have characteristics of both metals and nonmetals are called metalloids.

Atomic radius is defined as one half the distance between nuclei of two like atoms.

The amount of energy released when an atom gains an electron is called electron affinity.

The amount of energy required to remove an electron from a neutral atom is called ionization energy.

# Review Sheet Unit 4 Chemistry Answers

**Ying-Ying Zheng**



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