

Chapter 5 Study Guide- Physical Science
Test Monday, April 28th, 2014

1. What is a chemical bond? **force of attraction that holds elements together in order to form a unit**
2. Define chemical structure. **the arrangement of atoms in a substance.**
3. What is an ionic bond? **a chemical bond in which one atom loses an electron to form a positive ion and the other atom gains an electron to form a negative ion**
4. An ionic bond is between a **METAL** and a **NONMETAL** .
5. What is a covalent bond? **a chemical bond that involves sharing a pair of electrons between atoms in a molecule**
6. A covalent bond is between a **nonmetal** and a **nonmetal** .
7. What is a metallic bond? **a bond formed by the attraction between positively charged metal ions and the electrons around them.**
8. What are the polyatomic ions? **an ion made of two or more atoms.**
9. What is an ION? **a charged atom**
10. What is a CATION? **A positively charged atom**
11. What is an ANION? **A negatively charged atom**
12. Which part of an ionic compound should be written first? **Metal**

Name the following ionic compounds:

13. FeI₂- **Iron Iodide**
14. MnF₃- **Manganese Fluoride**
15. CrCl₂- **Chromium Chloride**
16. CuS- **Copper Sulfide**

Name the following covalent compounds:

17. As₂O₅- **diarsenic pentaoxide**
18. SiI₄- **silicon tetraiodide**
19. P₄S₃- **tertraphosphorus trisulfide**
20. P₄O₁₀- **tetraphosphorus decoxide**

Write the equation for the following ionic compounds:

21. Lithium Oxide- **Li₂O**
22. Magnesium Nitride- **Mg₃N₂**
23. Sodium Phosphide- **Na₃P**
24. Calcium Bromide -**CaBr₂**

Define/Explain the following

25. What is a subscript? **number written below, tells you how many of each element is present in compound/molecule**
26. What is a superscript? **number written above the element symbol, this is how the oxidation number (charge) is displayed.**

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R Sanford



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Table of Contents Physical Science Chemical Bond Study Guide

1. Understanding the eBook Physical Science Chemical Bond Study Guide
 - The Rise of Digital Reading Physical Science Chemical Bond Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Chemical Bond Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Chemical Bond Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Chemical Bond Study Guide
 - Personalized Recommendations
 - Physical Science Chemical Bond Study Guide User Reviews and Ratings
 - Physical Science Chemical Bond Study Guide and Bestseller Lists
5. Accessing Physical Science Chemical Bond Study Guide Free and Paid eBooks

- Physical Science Chemical Bond Study Guide Public Domain eBooks
- Physical Science Chemical Bond Study Guide eBook Subscription Services
- Physical Science Chemical Bond Study Guide Budget-Friendly Options
- 6. Navigating Physical Science Chemical Bond Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science Chemical Bond Study Guide Compatibility with Devices
 - Physical Science Chemical Bond Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Chemical Bond Study Guide
 - Highlighting and Note-Taking Physical Science Chemical Bond Study Guide
 - Interactive Elements Physical Science Chemical Bond Study Guide
- 8. Staying Engaged with Physical Science Chemical Bond Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Chemical Bond Study Guide
- 9. Balancing eBooks and Physical Books Physical Science Chemical Bond Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Chemical Bond Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science Chemical Bond Study Guide
 - Setting Reading Goals Physical Science Chemical Bond Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science Chemical Bond Study Guide
 - Fact-Checking eBook Content of Physical Science Chemical Bond Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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