

- 99) Ionization energies allow chemists to determine electron structures. The first two ionization energies of rubidium and strontium are given in the table.

element	first ionization energy (kJ/mol)	second ionization energy (kJ/mol)
rubidium	403	2633
strontium	550	1064

- 100) Write an equation to represent the **second** ionization energy of strontium.

Include state symbols.

[2]

- 101) Why is the **first** ionization energy of strontium larger than the **first** ionization energy of rubidium?

Explain answer (as always) use appropriate chemical terms **specifically**.

[2]

- 102) Why is the **second** ionization energy of rubidium larger than the **second** ionization energy of strontium?

[2]

(Total 12)

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Table of Contents Ocr Chemistry F321 Jan 13 Paper

1. Understanding the eBook Ocr Chemistry F321 Jan 13 Paper
 - The Rise of Digital Reading Ocr Chemistry F321 Jan 13 Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Ocr Chemistry F321 Jan 13 Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ocr Chemistry F321 Jan 13 Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr Chemistry F321 Jan 13 Paper
 - Personalized Recommendations
 - Ocr Chemistry F321 Jan 13 Paper User Reviews and Ratings
 - Ocr Chemistry F321 Jan 13 Paper and Bestseller Lists
5. Accessing Ocr Chemistry F321 Jan 13 Paper Free and Paid eBooks
 - Ocr Chemistry F321 Jan 13 Paper Public Domain eBooks
 - Ocr Chemistry F321 Jan 13 Paper eBook Subscription Services
 - Ocr Chemistry F321 Jan 13 Paper Budget-Friendly Options

6. Navigating Ocr Chemistry F321 Jan 13 Paper eBook Formats
 - ePub, PDF, MOBI, and More
 - Ocr Chemistry F321 Jan 13 Paper Compatibility with Devices
 - Ocr Chemistry F321 Jan 13 Paper Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ocr Chemistry F321 Jan 13 Paper
 - Highlighting and Note-Taking Ocr Chemistry F321 Jan 13 Paper
 - Interactive Elements Ocr Chemistry F321 Jan 13 Paper
8. Staying Engaged with Ocr Chemistry F321 Jan 13 Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ocr Chemistry F321 Jan 13 Paper
9. Balancing eBooks and Physical Books Ocr Chemistry F321 Jan 13 Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ocr Chemistry F321 Jan 13 Paper
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ocr Chemistry F321 Jan 13 Paper
 - Setting Reading Goals Ocr Chemistry F321 Jan 13 Paper
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ocr Chemistry F321 Jan 13 Paper
 - Fact-Checking eBook Content of Ocr Chemistry F321 Jan 13 Paper
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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