(a) Interpret and use the terms:

(i) Empirical formula = simplest whole number ratio of atoms of each element present in a compound

(ii) Molecular formula = actual number of atoms of each element in a molecule,

(iii) General formula = simplest algebraic formula of a member of a homologous series

| Alkane | CoH2n+2 |
|--------------|-------------------|
| Alkene | C-Ha- |
| Alcohol | C-H22OH |
| Cycloalkane | C-H _{2m} |
| Cycloalicene | Call Page 1 |
| Haloalkane . | C-HX |

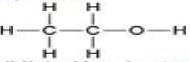
(iv) Structural formula = minimal detail that shows the arrangement of atoms in a molecule

E.g. Butane:

CH3CH2CH2CH3 or CH3(CH2)2CH3,

(v) Displayed formula = relative positioning of atoms and the bonds between them,

i.e. for ethanol:



(vi) Skeletal formula = simplified organic formula, shown by removing hydrogen atoms from alkyl chains, leaving just a carbon skeleton and associated functional group

ie for butan-1-ol:



(b) Interpret, and use, the terms:

(i) hamologous series = a series of organic compounds having the same functional group but with each successive member differing by CH2.

(ii) functional group = a group of atoms responsible for the characteristic reactions of a compound;

(c) Use the general formula of a homologous series to predict the formula of any member of the series;

E.g - Alkane = C_nH_{2n+2} If C = 9, $H = (9 \times 2) + 2$

Therefore = C_9H_{20} = Nonane

(d) State the names of the first ten members of the alkane homologous series;

| No of Carbons | Alkyl Group |
|---------------|-------------|
| | Methyl |
| 2 | Ethyl |
| 3 | Propyl |
| 4 | Butyl |
| 5 | Pentyl |
| 6 | Hiescyl |
| 7 | Heptyl |
| 8 | Octyl |
| 9 | Nonyl |
| 10 | Decyl |

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