



**Tuesday 21 May 2024 – Afternoon**

**GCSE (9–1) Computer Science**

**J277/02 Computational thinking, algorithms and programming**

**Time allowed: 1 hour 30 minutes**

Do not use:  
• a calculator



Please write clearly in black ink. Do not write in the barcodes.

Centre number

Candidate number

First name(s) \_\_\_\_\_

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**INSTRUCTIONS**

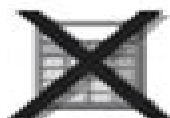
- Use black ink.
- Write your answer to each question in the space provided. If you need extra space use the lined pages at the end of this booklet. The question numbers must be clearly shown.
- Answer **all** the questions.

**INFORMATION**

- The total mark for this paper is **80**.
- The marks for each question are shown in brackets [ ].
- This document has **20** pages.

**ADVICE**

- Read each question carefully before you start your answer.
- We advise you to spend approximately 50 minutes on Section A and approximately 40 minutes on Section B.



# Ocr B711 2014 May Mark

**O García**



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Mark Coffey, Dennis Brown, 2016 Religion and Ethics for OCR is an ideal guide for students taking either the new AS level or A level qualifications offered by OCR Drawing on insights gained over years of teaching and following the OCR course outlines closely Mark Coffey and Dennis Brown's landmark book includes up to date discussions of key debates in religion and ethics ethical theories set in their historical context discussion of contemporary developments in science technology and society helpful guidance on writing the perfect exam answer and detailed guidance on OCR's assessment criteria and how best to meet them

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web basically there are five types of chemical equations and their reactions check them out below combination or synthesis chemical reaction this is the most common type of chemical equation in this chemical equation a new product is formed by combining two to three combinations of reactants for instance  $h_2 + o_2 \rightarrow h_2o$  this is a chemical



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 $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$   
 $\text{Mg}(\text{OH})_2 + 2\text{HNO}_3 \rightarrow \text{Mg}(\text{NO}_3)_2 + 2\text{H}_2\text{O}$   
 $2\text{NaClO}_3 \rightarrow 2\text{NaCl} + 3\text{O}_2$   
 $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$   
 $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$   
 $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$

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 $\text{Mg}(\text{OH})_2 + \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2\text{O}$   
 note all reactants and products require a coefficient of at least one  
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 $2\text{C}_2\text{H}_2 + 5\text{O}_2 \rightarrow 4\text{CO}_2 + 2\text{H}_2\text{O}$   
 if 1.0 mol of  $\text{C}_2\text{H}_2$  is burned how many moles of  $\text{CO}_2$  are formed  
 if 1.0 mol of  $\text{C}_2\text{H}_2$  is burned how many moles of oxygen gas are consumed

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 nitric acid  
 potassium hydroxide  
 water  
 zinc  
 zinc nitrate  
 copper

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