

### Chemical Formula

Talking about proportions molecules and relative of atoms present in a compound

Molecular  $\text{C}_{10}\text{H}_{18}\text{O}$  (molecular)

Formula represents the number of atoms of each  
in a single molecule

Empirical  $\text{C}_5\text{H}_9\text{O}$  (molecular)

Formula represents the simplest ratio of  
positive and negative ions in a formula  
unit

### Atomic Mass

The average mass of atoms present in a sample  
unit

### Mass

$m = \text{mass}$  unit mass

$m = \text{mass}$  unit mass

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$m = \text{mass}$  unit mass

$m = \text{mass}$  unit mass

# Note Taking Guide Episode 601 Physics

**Scott C. Dulebohn**



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