

c. Standard Enthalpy of Vaporization ΔH_{vap}°

→ The enthalpy change for the vaporization of 1 mole of liquid to a gas at the boiling point and standard pressure.

Example:

$$\Delta H_{vap}^\circ = \Delta H_{fg}^\circ \quad (\Delta H_{vap}^\circ = +44.0 \text{ kJ mol}^{-1})$$

d. Standard Enthalpy of Melting ΔH_{fus}°

→ The enthalpy change for melting 1 mol of a solid into a liquid at the melting point and standard pressure.

Example:

$$\Delta H_{fus}^\circ = \Delta H_{sf}^\circ \quad (\Delta H_{fus}^\circ = +6.01 \text{ kJ mol}^{-1})$$

e. Standard Neutralization Enthalpy ΔH_{neut}°

→ The enthalpy change in the neutralization of an acid (H⁺) by a base (OH⁻) forming 1 mole of water.

Example:

$$\Delta H_{neut}^\circ = \Delta H_{aH}^\circ + \Delta H_{bOH}^\circ = \Delta H_{aH}^\circ + \Delta H_{sf}^\circ \quad (\Delta H_{neut}^\circ = -57.3 \text{ kJ mol}^{-1})$$

f. Standard Enthalpy of Burning ΔH_{comb}°

→ The enthalpy change for burning 1 mol of elemental atoms in the gas phase at standard conditions.

Example:

$$\Delta H_{comb}^\circ = \Delta H_{fg}^\circ \quad (\Delta H_{comb}^\circ = -44.0 \text{ kJ mol}^{-1})$$

g. Standard Solution Enthalpy ΔH_{sol}°

→ The enthalpy change when 1 mole of a compound is converted into a solution at standard conditions.

Example:

$$\Delta H_{sol}^\circ = \Delta H_{aH}^\circ \quad (\Delta H_{sol}^\circ = -23 \text{ kJ mol}^{-1})$$

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Robert R. Redfield



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