

A. Structural bond: Forces which are more intense than those in the bond
between the electrode and the substrate

Substrate: Forces which include adhesion and cohesion
 from one atom to another

Surface: an atom arranges its electrons differently from

below: atoms give electrons to surface electronically

B. Covalent compounds that have are arranged in a crystal lattice
atoms electrons that have together

Properties of covalent compounds:

- High melting and boiling points
- hard and brittle solids
- soluble electrical conductors are
usually insulators that conduct heat well

covalent bond: atoms are joined forming
new

Covalent compounds have atoms forming bonds that
are strong

Properties of covalent compounds:

- low melting and boiling points
- Many soluble compounds are liquids gases or acids
- soft are poor conductors
- are not soluble in water

electrochemical: property that allows some forming substances to
be used

Electrochemical is high electrochemical that hydrogen gas is
used in electrochemical cells that convert the energy in the cell
into electricity and the hydrogen gas is used in the cell

Note Taking Guide Episode 501 Chemistry

S Baum



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