

Chemical Formula

Talking about proportions molecules and relative of atoms present in a compound

Molecular $\text{C}_{10}\text{H}_{12}\text{O}_4$ (molecular)

Formula represents the number of atoms of each
in a single molecule

Empirical $\text{C}_5\text{H}_6\text{O}_2$ (molecular)

Formula represents the simplest ratio of
positive and negative ions in a formula
unit

Atomic Mass

The average mass of atoms present in a sample
unit

Mass

$m = \text{mass}$ unit mass

$m = \text{mass}$ unit mass

Mass

$m = \text{mass}$ unit mass

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$m = \text{mass}$ unit mass

Note Taking Guide Episode 601 Answers Physics

Rachel Sandford



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