

A reaction in which the _____ can react to form the _____ is called a _____ reaction.



Chemical _____ occurs when the _____ in a _____ reaction form _____ at the same _____ that _____ form _____.

At equilibrium:

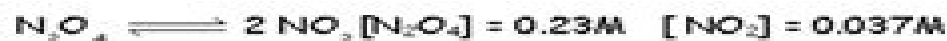
- the _____ of the reactants and products does not _____.
- the concentration of reactants can be _____ to, _____, or _____ the concentration of the products.



$$K_{eq} = \frac{[\text{_____}]^d [\text{_____}]^c}{[\text{_____}]^a [\text{_____}]^b}$$

If K_{eq} is	_____ reaction is favored
= 1	neither
< 1	
> 1	

Ex. #1: Calculate K_{eq} for the following reaction using the given equilibrium concentrations. Then determine whether the forward or reverse reaction is favored.



Note Taking Guide 1202

**Arabella Bowen, Todd Obolsky, Ross
Velton**



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