

What are Acids, Bases, and Salts?

The Swedish chemist Svante Arrhenius introduced the theory of ionization and used this theory to explain much about the behavior of acids and bases.

- An Arrhenius acid is defined as any compound that dissociates in aqueous solution to form H^+ ions.



- An Arrhenius base is defined as any compound that dissociates in aqueous solution to form OH^- ions.



- Salts are compounds that dissociate in aqueous solution releasing neither H^+ ions nor OH^- ions.



Using the **Arrhenius definition**, classify the following examples as acids, bases, or salts:

HBr	<u>acid</u>	KCl	<u>salt</u>
Mg(OH) ₂	<u>base</u>	H ₃ PO ₄	<u>acid</u>
HCl	<u>acid</u>	HClO	<u>acid</u>
KNO ₃	<u>salt</u>	Al(OH) ₃	<u>base</u>
HFO ₄	<u>acid</u>	KC ₂ H ₃ O ₂	<u>salt</u>
Ba(OH) ₂	<u>base</u>	NaCl	<u>salt</u>

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Basics of Analytical Chemistry and Chemical Equilibria Brian M. Tissue, 2013-07-22 Enables students to progressively build and apply new skills and knowledge Designed to be completed in one semester this text enables students to fully grasp and apply the core concepts of analytical chemistry and aqueous chemical equilibria Moreover the text enables readers to master common instrumental methods to perform a broad range of quantitative analyses Author Brian Tissue has written and structured the text so that readers progressively build their knowledge beginning with the most fundamental concepts and then continually applying these concepts as they advance to more sophisticated theories and applications Basics of Analytical Chemistry and Chemical Equilibria is clearly written and easy to follow with plenty of examples to help readers better understand both concepts and applications In addition there are several pedagogical features that enhance the learning experience including Emphasis on correct IUPAC terminology You Try It spreadsheets throughout the text challenging readers to apply their newfound knowledge and skills Online tutorials to build readers skills and assist them in working with the text s spreadsheets Links to analytical methods and instrument suppliers Figures illustrating principles of analytical chemistry and chemical equilibria End of chapter exercises Basics of Analytical Chemistry and Chemical Equilibria is written for undergraduate students who have completed a basic course in general chemistry In addition to chemistry students this text provides an essential foundation in analytical chemistry needed by students and practitioners in biochemistry environmental science chemical engineering materials science nutrition agriculture and the life sciences

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