Question	Answer	Marks	Guidance
1	Ca: N: O = 30,35/40.1: 21.20/14.0: 48.45/16.0 = 0.7569: 1.5286: 3.028	81	
	Formula = CaN <sub>2</sub> O <sub>4</sub>	B1	
2 (a)	2KClO <sub>3</sub> (s) → 2KCl(s) + 3O <sub>2</sub> (g)	B1	
2 (b)	4.50 × 10 <sup>-3</sup> mol	B1	
2 (c)	$n(KCIO_3) = 3.0 \times 10^{-3} \text{ mol}$	B1	
	$M(KCIO_3) = 122.6 \text{ g mol}^{-1}$	B1	
	mass of KCIO <sub>3</sub> = 0.3678 g	B1	
3	M(ZnSO <sub>4</sub> ) = 161.5 g mol <sup>-1</sup>	B1	
	$n(ZnSO_4) = 6.57 \times 10^{-3} \text{ mol}$	B1	
	$n(H_2O) = (1.893 - 1.061)/18 = 4.62 \times 10^{-2} \text{ mol}$ $x = n(H_2O)/n(ZnSO_4) = 7$	B1	
	Formula = ZnSO <sub>4</sub> •7H <sub>2</sub> O	B1	
4	$M(Na_2CO_3) = 106.0 \text{ g mol}^{-1}$	B1	
	Actual n(Na <sub>2</sub> CO <sub>3</sub> ) = 0.0234 mol	B1	
	Theoretical ri(Na <sub>2</sub> CO <sub>3</sub> ) = 0.0234 × 100/65 = 0.0360 mol	B1	
	Theoretical n(NaHCO <sub>3</sub> ) = 0.0720 mol	B1	
	Mass of NaHCO <sub>3</sub> = 0.0720 × 84.0 = 6.05 g	B1	
5 (a)	$Fe_2O_3(s) + 3CO(g) \rightarrow 2Fe(s) + 3CO_2(g)$	B1	
5 (b)	$M(Fe_2O_3) = 159.6 \text{ g mol}^{-1}$	B1	
	n(Fe <sub>2</sub> O <sub>3</sub> ) = (10 000 × 10 <sup>6</sup> )159.6 = 6.266 × 10 <sup>7</sup> mol	B1	
	n(Fe) = 1.253 × 10 <sup>8</sup> mol = 6.992 × 10 <sup>9</sup> g (6992 tonne)	B1	
6 (a)	Na : N : O = 27.1/23.0 : 16.5/14.0 : 56.4/16.0 = 1.178 : 1.179 : 3.525	B1	
	Formula = NaNO <sub>3</sub>	B1	
6 (b)	2NaNO₂(s) → 2NaNO₂(s) + O₂(g)	B1	

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**R Barnett** 

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