

| Question number | Answer                                                                                                                                                                                                                                                                                                                                  | Marks                      | Guidance |
|-----------------|-----------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|----------------------------|----------|
| 1               | $\text{Ca} : \text{N} : \text{O} = 30.35/40.1 : 21.20/14.0 : 48.45/16.0$<br>$= 0.7569 : 1.5286 : 3.028$<br>Formula = $\text{CaN}_2\text{O}_4$                                                                                                                                                                                           | B1<br><br>B1               |          |
| 2 (a)           | $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$                                                                                                                                                                                                                                                    | B1                         |          |
| 2 (b)           | $4.50 \times 10^{-3} \text{ mol}$                                                                                                                                                                                                                                                                                                       | B1                         |          |
| 2 (c)           | $n(\text{KClO}_3) = 3.0 \times 10^{-3} \text{ mol}$<br>$M(\text{KClO}_3) = 122.6 \text{ g mol}^{-1}$<br>mass of $\text{KClO}_3 = 0.3678 \text{ g}$                                                                                                                                                                                      | B1<br>B1<br>B1             |          |
| 3               | $M(\text{ZnSO}_4) = 161.5 \text{ g mol}^{-1}$<br>$n(\text{ZnSO}_4) = 6.57 \times 10^{-3} \text{ mol}$<br>$n(\text{H}_2\text{O}) = (1.893 - 1.061)/18 = 4.62 \times 10^{-2} \text{ mol}$<br>$x = n(\text{H}_2\text{O})/n(\text{ZnSO}_4) = 7$<br>Formula = $\text{ZnSO}_4 \cdot 7\text{H}_2\text{O}$                                      | B1<br>B1<br>B1<br>B1       |          |
| 4               | $M(\text{Na}_2\text{CO}_3) = 106.0 \text{ g mol}^{-1}$<br>Actual $n(\text{Na}_2\text{CO}_3) = 0.0234 \text{ mol}$<br>Theoretical $n(\text{Na}_2\text{CO}_3) = 0.0234 \times 100/65$<br>$= 0.0360 \text{ mol}$<br>Theoretical $n(\text{NaHCO}_3) = 0.0720 \text{ mol}$<br>Mass of $\text{NaHCO}_3 = 0.0720 \times 84.0 = 6.05 \text{ g}$ | B1<br>B1<br>B1<br>B1<br>B1 |          |
| 5 (a)           | $\text{Fe}_2\text{O}_3(\text{s}) + 3\text{CO}(\text{g}) \rightarrow 2\text{Fe}(\text{s}) + 3\text{CO}_2(\text{g})$                                                                                                                                                                                                                      | B1                         |          |
| 5 (b)           | $M(\text{Fe}_2\text{O}_3) = 159.6 \text{ g mol}^{-1}$<br>$n(\text{Fe}_2\text{O}_3) = (10\,000 \times 10^6)/159.6 = 6.266 \times 10^7 \text{ mol}$<br>$n(\text{Fe}) = 1.253 \times 10^8 \text{ mol} = 6.992 \times 10^5 \text{ g}$<br>(6992 tonnes)                                                                                      | B1<br>B1<br>B1             |          |
| 6 (a)           | $\text{Na} : \text{N} : \text{O} = 27.1/23.0 : 16.5/14.0 : 56.4/16.0$<br>$= 1.178 : 1.179 : 3.525$<br>Formula = $\text{NaNO}_3$                                                                                                                                                                                                         | B1<br><br>B1               |          |
| 6 (b)           | $2\text{NaNO}_3(\text{s}) \rightarrow 2\text{NaNO}_2(\text{s}) + \text{O}_2(\text{g})$                                                                                                                                                                                                                                                  | B1                         |          |

# Ocr Chemistry Exam Question Answers

**F Rizvi**



## Ocr Chemistry Exam Question Answers:

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 Written by a senior examiner Mike Smith this OCR A AS Chemistry Student Unit Guide is the essential study companion for Unit F322 Chains Energy and Resources This full colour book includes all you need to know to prepare for your unit exam clear guidance on the content of the unit with topic summaries knowledge check questions and a quick reference index examiner s advice throughout so you will know what to expect in the exam and will be able to demonstrate the skills required exam style questions with graded student responses so you can see clearly what is required to get a better grade **Revise AS Chemistry for Salters (OCR)** Ann Daniels, 2005-02-22 Help students pull together the chemical ideas in the course and apply them to fresh contexts in exam questions *OCR(A) AS Chemistry Student Unit Guide: Unit F321 Atoms, Bonds and Groups* Mike Smith, 2012-01-13 AS Chemistry Student Unit Guide is the essential study companion for Unit F321 Atoms Bonds and Groups This book includes all you need to know to prepare for your unit exam clear guidance on the content of the unit with topic summaries knowledge check questions and a quick reference index examiner s advice throughout so you will know what to expect in the exam and will be able to demonstrate the skills required and exam style questions with graded student responses so you can see clearly what is required to get a better grade **My Revision Notes: OCR (A) A2 Chemistry** Mike Smith, 2012-05-18 Get the best grades with My Revision Notes OCR A A2 Chemistry Manage your own revision with step by step support from senior examiner and author Mike Smith Use specific examples and advice to improve your knowledge of chemical applications and processes Get the top marks by applying chemical terms accurately with the help of definitions and key words Improve your exam skills with the help of self testing and exam style questions and answers My Revision Notes will help you prepare for the big day Plan and pace your revision with My Revision Planner Use the concise notes to revise the essential information Use the examiner s tips and summaries to clarify key points Avoid making typical mistakes with expert advice Test yourself with end of topic questions and answers and tick off each topic as you complete it Practise your exam skills on exam questions then check your answers online Get exam ready with last minute quick quizzes at [www.therevisionbutton.co.uk/myrevisionnotes](http://www.therevisionbutton.co.uk/myrevisionnotes) *OCR A Chemistry A2 Student Unit Guide: Unit F325 New Edition: Equilibria, Energetics and Elements* ePub Mike Smith, 2013-02-22 Written by a former senior examiner Mike Smith this OCR A A2 Chemistry Student Unit Guide is the essential study companion for Unit F325 Equilibria Energetics and Elements This full colour book includes all you need to know to prepare for your unit exam clear guidance on the content of the unit with topic summaries knowledge check questions and a quick reference index examiner s advice throughout so you will know what to expect in the exam and will be able to demonstrate the skills required exam style questions with graded student responses so you can see clearly what is required to get a better grade [The Essentials of GCSE OCR Science for Specification A](#), 2006 *My Revision Notes: OCR (A) AS Chemistry* Mike Smith, 2012-04-20 Get the best grades with My Revision Notes OCR A AS Chemistry Manage your own revision with step by step support from senior examiner and author

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**Units F323 & F326 Practical Skills in Chemistry** John Older,2012-10-26 Written by a senior examiner John Older this OCR AS A2 Chemistry Student Unit Guide is the essential study companion for Units F323 and F326 Practical Skills in Chemistry This full colour book includes all you need to know to prepare for your Unit F323 and Unit F326 assessments clear guidance on the range of practical apparatus and techniques that you need to know about and an overview of the scientific method of testing ideas by experimentation examiner s advice throughout so you will know what to expect in the assessments and will be able to demonstrate the skills required sample investigation tasks for extra practice before your assessments

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### Table of Contents **Ocr Chemistry Exam Question Answers**

1. Understanding the eBook **Ocr Chemistry Exam Question Answers**
  - The Rise of Digital Reading **Ocr Chemistry Exam Question Answers**
  - Advantages of eBooks Over Traditional Books
2. Identifying **Ocr Chemistry Exam Question Answers**
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an **Ocr Chemistry Exam Question Answers**
  - User-Friendly Interface
4. Exploring eBook Recommendations from **Ocr Chemistry Exam Question Answers**
  - Personalized Recommendations
  - **Ocr Chemistry Exam Question Answers** User Reviews and Ratings
  - **Ocr Chemistry Exam Question Answers** and Bestseller Lists



5. Accessing Ocr Chemistry Exam Question Answers Free and Paid eBooks
  - Ocr Chemistry Exam Question Answers Public Domain eBooks
  - Ocr Chemistry Exam Question Answers eBook Subscription Services
  - Ocr Chemistry Exam Question Answers Budget-Friendly Options
6. Navigating Ocr Chemistry Exam Question Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Ocr Chemistry Exam Question Answers Compatibility with Devices
  - Ocr Chemistry Exam Question Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ocr Chemistry Exam Question Answers
  - Highlighting and Note-Taking Ocr Chemistry Exam Question Answers
  - Interactive Elements Ocr Chemistry Exam Question Answers
8. Staying Engaged with Ocr Chemistry Exam Question Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ocr Chemistry Exam Question Answers
9. Balancing eBooks and Physical Books Ocr Chemistry Exam Question Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ocr Chemistry Exam Question Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ocr Chemistry Exam Question Answers
  - Setting Reading Goals Ocr Chemistry Exam Question Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ocr Chemistry Exam Question Answers
  - Fact-Checking eBook Content of Ocr Chemistry Exam Question Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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