

Question	Answer	Marks	Guidance								
1 (a) (i)	<p>Atom(s) of an element</p> <p><b>AND</b></p> <p>with different numbers of neutrons (and with different masses) ✓</p>	1	<p><b>ALLOW</b> for 'atoms of an element':</p> <p>Atoms of the same element</p> <p>OR atoms with the same number of protons</p> <p>OR atoms with the same atomic number</p> <p><b>IGNORE</b> 'different relative atomic masses'</p> <p><b>IGNORE</b> different mass number</p> <p><b>IGNORE</b> same number of electrons</p> <p><b>DO NOT ALLOW</b> different numbers of electrons</p> <p><b>DO NOT ALLOW</b> 'atoms of elements' for 'atoms of an element'</p> <p><b>DO NOT ALLOW</b> 'an element with different numbers of neutrons' (ie atom(s) is essential)</p>								
	<p>(ii)</p> <table border="1" data-bbox="208 611 846 685"> <tr> <td>Protons</td><td>Neutrons</td><td>Electrons</td><td></td></tr> <tr> <td>74</td><td>110</td><td>74</td><td>✓</td></tr> </table>	Protons	Neutrons	Electrons		74	110	74	✓	1	
Protons	Neutrons	Electrons									
74	110	74	✓								
	(iii) $^{12}\text{C}$ OR C-12 OR carbon 12 OR carbon-12 ✓	1	<b>IGNORE</b> $1/12^{\text{th}}$ <b>AND</b> amu								
(b)	<p>(i)</p> <p>(Oxidised):</p> <p>H (oxidation number has increased) from H = 0 to H = +1 ✓</p> <p>(Reduced):</p> <p>W (oxidation number has decreased) from W = +6 to W = 0 ✓</p>	2	<p><b>ALLOW</b> 6+ OR 6 OR 1+ OR 1</p> <p><b>ALLOW</b> one mark for correct oxidation number changes</p> <p>H = 0 to H = +1 <b>AND</b> W = +6 to W = 0</p> <p><b>ALLOW</b> oxidation states written above the equation if not seen in the text <b>BUT IGNORE</b> oxidation states written above the equation if seen in the text</p> <p><b>ALLOW</b> for one mark: (Oxidised) H has increased by 1 <b>AND</b> (Reduced) W has decreased by 6</p> <p><b>IGNORE</b> <math>\text{WO}_3</math> is reduced</p> <p><b>IGNORE</b> references to electron loss / gain if correct</p> <p><b>DO NOT ALLOW</b> incorrect references to electron loss / gain</p> <p><b>DO NOT ALLOW</b> 'H oxidised and W reduced' without reference to oxidation number changes</p>								

# Ocr F321 Jan 2013 Question Paper

**M Planty**



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## **Table of Contents Ocr F321 Jan 2013 Question Paper**

1. Understanding the eBook Ocr F321 Jan 2013 Question Paper
  - The Rise of Digital Reading Ocr F321 Jan 2013 Question Paper
  - Advantages of eBooks Over Traditional Books
2. Identifying Ocr F321 Jan 2013 Question Paper
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ocr F321 Jan 2013 Question Paper
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr F321 Jan 2013 Question Paper
  - Personalized Recommendations
  - Ocr F321 Jan 2013 Question Paper User Reviews and Ratings
  - Ocr F321 Jan 2013 Question Paper and Bestseller Lists
5. Accessing Ocr F321 Jan 2013 Question Paper Free and Paid eBooks
  - Ocr F321 Jan 2013 Question Paper Public Domain eBooks
  - Ocr F321 Jan 2013 Question Paper eBook Subscription Services
  - Ocr F321 Jan 2013 Question Paper Budget-Friendly Options

6. Navigating Ocr F321 Jan 2013 Question Paper eBook Formats
  - ePub, PDF, MOBI, and More
  - Ocr F321 Jan 2013 Question Paper Compatibility with Devices
  - Ocr F321 Jan 2013 Question Paper Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ocr F321 Jan 2013 Question Paper
  - Highlighting and Note-Taking Ocr F321 Jan 2013 Question Paper
  - Interactive Elements Ocr F321 Jan 2013 Question Paper
8. Staying Engaged with Ocr F321 Jan 2013 Question Paper
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ocr F321 Jan 2013 Question Paper
9. Balancing eBooks and Physical Books Ocr F321 Jan 2013 Question Paper
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ocr F321 Jan 2013 Question Paper
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ocr F321 Jan 2013 Question Paper
  - Setting Reading Goals Ocr F321 Jan 2013 Question Paper
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ocr F321 Jan 2013 Question Paper
  - Fact-Checking eBook Content of Ocr F321 Jan 2013 Question Paper
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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