

of substance:

Substance:
 mass of particles of a substance
 in mol

6.02×10^{23} particles
 amount of substance

Constant:
 particles per mol, $6.02 \times 10^{23} \text{ mol}^{-1}$

mol, measured in g
 equivalent to M_r

Molar Gas Volume:
 volume of gas per mole of gas molecules in a stated temperature + pressure
 - Measured in $\text{dm}^3 \text{ mol}^{-1}$ or $\text{m}^3 \text{ mol}^{-1}$
 At RTP (20°C and 100kPa):
 (293K and 100000 Pa)
 is $24 \text{ dm}^3 \text{ mol}^{-1}$
 at RTP.

Volume /
 dm^3

Electronegativity: → the measure of attraction of a bonded atom in a covalent bond

ELECTRONEGATIVITY INCREASES ACROSS PERIOD:
 - Increase in no. of protons in an atom
 → Nuclear charge increases

ELECTRONEGATIVITY DE:
 - More electron shells
 → More shielding
 atomic radius increases
 nuclear charge
 and nuclear attraction
 is less electronegative.

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Atomic Structure

Electrons:
 → shared in the same energy level
 of p and s but different
 shells different sizes

Physical Properties:
 → can be different - e.g. different
 melting + boiling points
 → i.e. massy water has higher melting
 point + boiling point + density

Chemical Properties:
 → same no. of electrons + same electronic
 structure in outer shell + transfer of e⁻ is
 unaffected hence have same chemical
 properties.
 → different no. of e⁻ has no effect
 on bonding + reactions



→ consists of a small dense
 nucleus of protons + neutrons
 surrounded by electrons orbiting
 in electron orbitals in shells;
 - No. of p⁺ = No. of e⁻
 → Equal + opposite charges cancel
 so atom has neutral charge
 - Neutrons important in nuclear
 stability - electrostatic repulsion
 between p⁺ in small dense nucleus

Mass Number:
 → No. of p⁺ + n in nucleus of atom
 - p⁺ and n have relative mass of 1
 and e⁻ mass is negligible to 0

Relative Isotopic Mass:
 → the mass of an atom of an isotope
 relative to 1/12th the mass of an atom of C-12

Relative Atomic Mass:
 → the weighted mean mass of an atom of
 an element relative to 1/12th the mass of
 an atom of C-12

Weighted Mean Mass:
 → average mass of all isotopes, taking into
 account the relative abundance of each isotope

- If dipole
 moment is
 out + molts

e.g. H₂O

- O is more electronegative than H
 - O-H has permanent dipole
 hence do not cancel out
 there is an overall dipole

e.g. CO₂

O is more electronegative than C
 C=O has a dipole
 but the bond
 dipoles cancel out

Compounds + ions Formulae:

Alkali metal Halogenoalkane HX_n

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Mark Steyvers



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