(d) The student found that 22.50 cm³ of 0.0800 moldm⁻³ Na₂B₄O₇ reacted with 25.00 cm³ of dilute hydrochloric acid.

$$Na_2B_4O_7 + 2HCl + 5H_2O \rightarrow 2NaCl + 4H_3BO_3$$

(i) Calculate the amount, in mol, of Na₂B₄O₇ used.

(ii) Calculate the amount, in mol, of HCl used.

(iii) Calculate the concentration, in moldm⁻³, of the HCL

[Total: 12]

END OF QUESTION PAPER



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