

Nursing Case Study

I. Health History

The subject of my case study is W. K. W.K. is a 95 year old white male who was admitted to Moses Cone Hospital on November 23, 2002. Prior to being admitted to the hospital, W. K. had been in excellent health. His troubles apparently began three weeks prior to being admitted. On November 23, W. K.'s son found him lying on the floor confused, and soaked in urine.

Mr. K. was diagnosed as having an acute cerebral vascular accident. This disorder can also be described as a "stroke". It occurs when there is an interruption of normal blood flow in one or more of the blood vessels that supply the brain. Thrombosis, embolism, and hemorrhage are the primary causes of a CVA. (Sommers and Johnson 2002) The tissues of the brain become ischemic, leading to hypoxia or anoxia with destruction or necrosis of the neurons, glia, and vasculature. Complications of CVA include unstable blood pressure, sensory and motor impairment, infection, pneumonia, contractures, and pulmonary emboli. CVA is the third leading cause of death in the United States and affects more than 500,000 Americans annually. (Sommers and Johnson 2002)

He was widowed in October of 2001, one daughter has coronary artery disease, one son died of an MI at age 37, and one son died with lung cancer at 57. He had been the primary care giver of his daughter until she was admitted to the hospital three weeks ago. She is dying with a short bowel syndrome and cirrhosis and is now being taken care of in hospice. Soon after being left alone, Mr. K.'s appetite decreased and he had become congested. He was placed on Paxil to treat symptoms of depression. He had also been taking Cipro for congestion. Also his family noticed that he was suffering from confusion. As a result, they brought him into the Emergency Room for evaluation. The Emergency Room doctors performed a CT scan of the brain which revealed evidence of old strokes. The doctors

Sample Case Studies Nursing

Leslie Neal-Boylan



Sample Case Studies Nursing:

The Enigmatic Realm of **Sample Case Studies Nursing**: Unleashing the Language is Inner Magic

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web objective to apply stoichiometric principles to calculate the appropriate chemical equation for the reaction between iron metal and a solution of copper i sulfate introduction in the form of finely split powder copper metal is produced as a byproduct of the process as iron ions the iron metal dissolves

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web purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be conducted this will help you to differentiate limiting and excess reactant in a chemical reaction finally the theoretical and percent yield of this reaction will be calculated

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web purpose in this experiment you will observe a single replacement reaction you will also use stoichiometry to predict what the theoretical yield of product is and calculate a percent yield reaction iron copper ii sulfate copper iron ii sulfate pre lab questions write the balanced equation

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web october 3 2017 abstract the techniques used in this lab are quantitative transfer and vacuum filtration with the reaction of 8 001 grams of copper ii sulfate CuSO_4 and 2 0153 grams of iron powder Fe the goal of this experiment was to determine the product of copper ii sulfate with iron

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4

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web stoichiometry lab the reaction of iron with copper ii sulfate the study of stoichiometry deals with the calculation of quantities in a chemical reaction how much product will be produced how much reactant do you need to make that much product

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web description great lab for use during stoichiometry unit in chemistry reinforces ideas of mole ratios and stoichiometry calculations by measuring the mass of iron that reacts and the mass of copper metal produced students can calculate the ratio of moles of reactant to moles of product this experimental mole ratio from lab can be compared

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper ii sulfate this reaction produces metallic copper which is seen precipitating as a finely divided red powder

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web in this lab we will use experimental evidence and the principles of stoichiometry to determine which ion is formed in a single replacement reaction between iron and copper ii sulfate pre lab assignment in preparation for the lab activity we will investigate another single replacement reaction between copper and silver nitrate

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of

substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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web stoichiometric principles can be used to determine the chemical equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction the iron ions dissolve in the copper ii sulfate solution to produce copper ii ions

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web under the finalize of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away copper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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web stoichiometry lab iron and copper sulfate sample data data and observations mass of copper ii sulfate pentahydrate 3 27 g of g mass of filter paper mass of filter paper and dry copper mass of dry copper 1 28 g

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