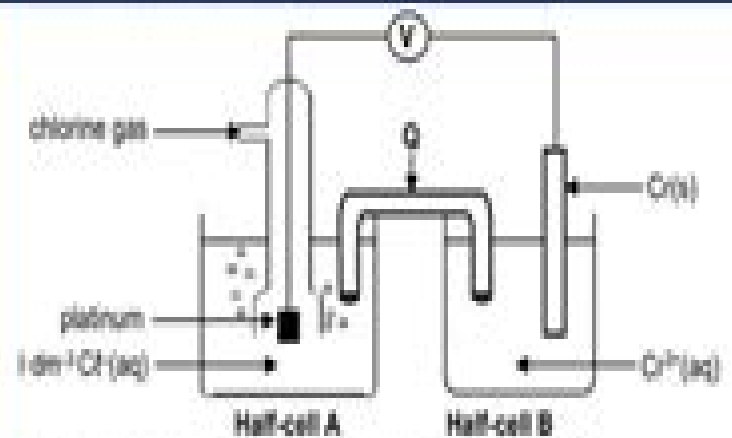
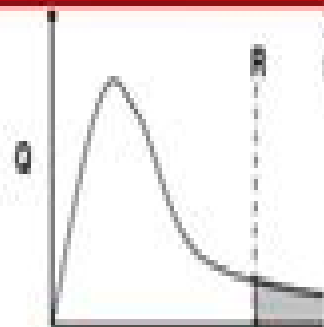


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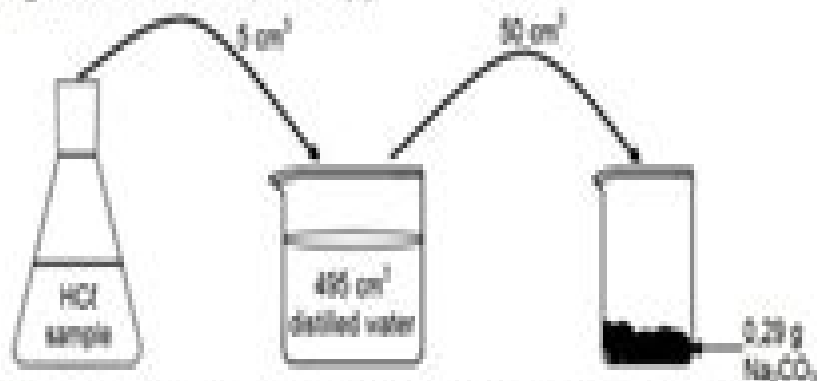


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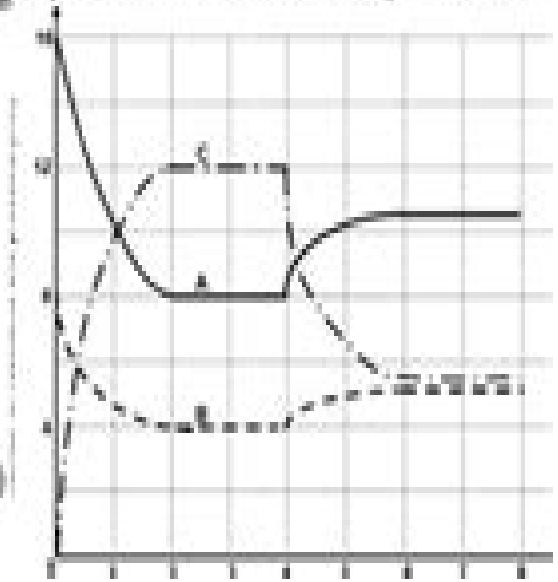
LIVE

During a reaction 50 cm³ of this dilute hydrochloric acid solution, HCl(aq) , reacts completely with 0.29 g of sodium carbonate, $\text{Na}_2\text{CO}_3(\text{s})$.



Write a balanced equation for the reaction (s): $\text{Na}_2\text{CO}_3(\text{s}) + 2\text{HCl(aq)} \rightarrow 2\text{NaCl(aq)} + \text{CO}_2(\text{g}) + \text{H}_2\text{O(l)}$
 Calculate the concentration of the hydrochloric acid sample.
 Answer: 10.94 mol/dm³

Graph of number of moles of reactants and products versus time



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