



MERCY JUNIOR COLLEGE
Grade 11
Physical Science
Finals Exam

NAME: _____

YEAR LEVEL & SECTION: _____

TEACHER'S NAME: _____

PROCTOR'S NAME: _____

1. MULTIPLE CHOICE

INSTRUCTION: Encircle the letter of the correct answer. (1 POINT EACH)

- It is a subatomic particle, symbol e^- or β^- , with a negative elementary electric charge.
A. Neutron
B. Electron
C. Proton
D. Atom
- It is a subatomic particle, symbol p or p^+ , with a positive electric charge of $+1e$ elementary charge and mass slightly less than that of a neutron.
A. Neutron
B. Electron
C. Proton
D. Atom
- It is a subatomic particle, symbol n or n^0 , with no net electric charge and a mass slightly larger than that of a proton.
A. Neutron
B. Electron
C. Proton
D. Atom
- Is this an example of ionic bonding?



carbon dioxide

- Yes, it is ionic bonding.
- No, it is of another type of chemical bonding.
- There is no chemical bonding present.
- This is just an orbital model.

- Is this an example of ionic bonding?



lithium fluoride

- Yes, it is ionic bonding.
 - No, it is of another type of chemical bonding.
 - There is no chemical bonding present.
 - This is just an orbital model.
- In atomic physics and quantum chemistry, the electron configuration is the distribution of electrons of an atom or molecule in atomic or molecular orbitals.
A. Electron Configuration
B. Periodic Table of Elements
C. Orbital Models
D. Energy Levels
 - Is this an example of ionic bonding?
-
- Yes, it is ionic bonding.
 - No, it is of another type of chemical bonding.
 - There is no chemical bonding present.
 - This is just an orbital model.
- It is a chemical rule of thumb that reflects observation that atoms of main-group elements tend to combine in such a way that each atom has eight electrons in its valence shell, giving it the same electronic configuration as a noble gas.
A. Octet Rule
B. Valence Electrons
C. Atomic Orbitals
D. Chemical Bonds
 - It is the smallest constituent unit of ordinary matter that has the properties of a chemical element.
A. Neutron
B. Electron
C. Proton
D. Atom
 - It is a type of chemical bond that involves the electrostatic attraction between oppositely

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Scope

N Noddings



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Arguing with Numbers James Wynn,G. Mitchell Reyes,2021-04-15 As discrete fields of inquiry rhetoric and mathematics have long been considered antithetical to each other That is if mathematics explains or describes the phenomena it studies with certainty persuasion is not needed This volume calls into question the view that mathematics is free of rhetoric Through nine studies of the intersections between these two disciplines Arguing with Numbers shows that mathematics is in fact deeply rhetorical Using rhetoric as a lens to analyze mathematically based arguments in public policy political and economic theory and even literature the essays in this volume reveal how mathematics influences the values and beliefs with which we assess the world and make decisions and how our worldviews influence the kinds of mathematical instruments we construct and accept In addition contributors examine how concepts of rhetoric such as analogy and visuality have been employed in mathematical and scientific reasoning including in the theorems of mathematical physicists and the geometrical diagramming of natural scientists Challenging academic orthodoxy these scholars reject a math equals truth reduction in favor of a more constructivist theory of mathematics as dynamic evolving and powerfully persuasive By bringing these disparate lines of inquiry into conversation with one another Arguing with Numbers provides inspiration to students established scholars and anyone inside or outside rhetorical studies who might be interested in exploring the intersections between the two disciplines In addition to the editors the contributors to this volume are Catherine Chaput Crystal Broch Colombini Nathan Crick Michael Dreher Jeanne Fahnestock Andrew C Jones Joseph Little and Edward Schiappa

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New Statesman ,1960 *Focus on Physical Science* , Study and Master Physical Science Grade 11 and 12 E. J. Van Zyl,V. Craul,A. Meyer,C. Muller,L. P. Spies,1999-01 *Everything Science* , **Physical Sciences 11** ,2012 *Core Objectives Mid-year and Final Examinations* Detroit Public Schools,1983

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Table of Contents Physical Science Grade 11 Final Examination 2014 Scope

1. Understanding the eBook Physical Science Grade 11 Final Examination 2014 Scope
 - The Rise of Digital Reading Physical Science Grade 11 Final Examination 2014 Scope
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Grade 11 Final Examination 2014 Scope
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Grade 11 Final Examination 2014 Scope
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Grade 11 Final Examination 2014 Scope
 - Personalized Recommendations
 - Physical Science Grade 11 Final Examination 2014 Scope User Reviews and Ratings
 - Physical Science Grade 11 Final Examination 2014 Scope and Bestseller Lists
5. Accessing Physical Science Grade 11 Final Examination 2014 Scope Free and Paid eBooks
 - Physical Science Grade 11 Final Examination 2014 Scope Public Domain eBooks
 - Physical Science Grade 11 Final Examination 2014 Scope eBook Subscription Services
 - Physical Science Grade 11 Final Examination 2014 Scope Budget-Friendly Options

6. Navigating Physical Science Grade 11 Final Examination 2014 Scope eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science Grade 11 Final Examination 2014 Scope Compatibility with Devices
 - Physical Science Grade 11 Final Examination 2014 Scope Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Grade 11 Final Examination 2014 Scope
 - Highlighting and Note-Taking Physical Science Grade 11 Final Examination 2014 Scope
 - Interactive Elements Physical Science Grade 11 Final Examination 2014 Scope
8. Staying Engaged with Physical Science Grade 11 Final Examination 2014 Scope
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Grade 11 Final Examination 2014 Scope
9. Balancing eBooks and Physical Books Physical Science Grade 11 Final Examination 2014 Scope
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Grade 11 Final Examination 2014 Scope
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Grade 11 Final Examination 2014 Scope
 - Setting Reading Goals Physical Science Grade 11 Final Examination 2014 Scope
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science Grade 11 Final Examination 2014 Scope
 - Fact-Checking eBook Content of Physical Science Grade 11 Final Examination 2014 Scope
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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