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Understanding Concepts

- 11. What is a stable electron configuration?
- 12. What does each dot in an electron dot disconsminepresent?
- 1.3. What process changes atoms into ions?
- 14. What keeps the ions in their fixed positions within a crystal lattice?
- 15. What are subscripts used for in chemical formulas?
- 16. Explain why a melted ionic compound is a good. conductor of electric current, but a solid ionic compound is a poor conductor of electric current.
- 1.7. What distinguishes single, double, and triple covalient bonds?
- Explain why the covalent bonds in molecules of July elements are always nonpolar.
- 19. Explain why, in a covalent bond between oxygen and hydrogen, the hydrogen atom has a partial positive charge and the oxygen atom has a partial megative charge.
- 20. What is the name of the binary compound. formed from potassium and iodine?
- 21. Write the formulas for the compounds called copper(I) chloride and copper(II) chloride.
- 22. Name the compounds represented by the spacefilling models labeled A, B, and C.



Petersont Silluteren



■ Oxygen ■ Carbon III Mitropen

- 23, in general, what determines the strength of
- 24. What properties of copper and tin change when these metals are mixed together to form bronze?
- 25. What advantage of magnesium is retained in magnesium alloys? What disadvantage is reduced?

Chemical Bonds 187

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11-16, 26, 30, 33 8 17-19, 27, 34-35, 39-40, 42 20-22, 28-29, 31-32, 36-38

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Assessment



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Reviewing Content

1. b 2. € 3. . . 4. di 5. C 45. In 7. d 8. 6. 9. c

Understanding Concepts

- In a stable electron configuration. the highest occupied energy level is filled with electrons.
- Each dot represents a valence. electron.
- 13. The transfer of electrons
- Attractions between neighboring. cations and anions keep the ions in fixed positions within the lattice.
- 15. A subscript is used to show the number of atoms of an element in a molecule or the ratio of ions in a crystal.
- When an ionic compound melts, ions can move away from their fixed locations in the crystal lattice.
- 17. Two atoms share two electrons in a single bond, four in a double bond, and six in a triple bond.
- 18. The covalent bonds in molecules of elements are always nonpolar because the atoms have the same attraction for electrons.
- 19. The oxygen atom has a greater attraction for electrons than the hydrogen atom does.
- 20. Potassium iodide
- 21. CuCl and CuCls
- 22. A is sulfur trioxide, B is carbon monoxide, and C is nitrogen dioxide
- 23. In general, the more valence electrons a metal has, the stronger the metal bonds are.
- 24. A mixture of copper and tin is harder and stronger than either metal in its pure form.
- 25. The advantage that is retained is that magnesium is a lightweight metal. The disadvantage that is reduced is that magnesium is a soft metal.

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Judd E. Hollander

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