F321 Revision of definitions to be learnt for exam

Isotopes	Atoms of the same element with different numbers of neutrons and different masses.
Relative atomic mass, Ar	The weighted mean mass of an atom of an element compared with one-twelfth of the mass of an atom of ^{12}C
Relative isotopic mass	The mass of an atom of an isotope compared with one-twelfth of the mass of an atom of carbon-12.
Relative formula mass	The relative formula mass of a compound is equal to the sum of the individual relative atomic masses.
Relative molecular mass, Mr	The weighted mean mass of a molecule compared with one-twelfth of the mass of an atom of carbon-12.
Mole	The amount of any substance containing as many particles as there are carbon atoms in exactly 12 g of the carbon-12 isotope.
Molecular formula	The number of atoms of each element in a molecule.
Empirical formula	The simplest whole-number ratio of atoms of each element present in a compound.
Acid	A species that is a proton donor.
Base	A species that is a proton acceptor.
Concentrated Solution	Solutions containing a higher number of moles of solute in a given volume of solution
Dilute Solution	Solutions containing a fewer number of moles of solute in a given volume of solution
Hydrated	Salt containing water molecules known as "water of crystallisation"
Anhydrous	A substance that contains no water of crystallisation
Oxidation	Loss of electrons, or an increase in oxidation number.
Reduction	Gain of electrons, or a decrease in oxidation number.
Redox reaction	A reaction in which both reduction and oxidation take place.
(First) ionisation energy	The energy required to remove one electron from each atom in one mole of gaseous atoms to form one mole of gaseous 1+ ions.
(Second) ionisation energy	The energy required to remove one electron from each ion in one mole of gaseous 1+ ions to form one mole of gaseous 2+ ions.

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