

Question number	Answer	Marks	Guidance
6 (c)	$M(\text{NaNO}_3) = 85.0 \text{ g mol}^{-1}$ $n(\text{NaNO}_3) = 0.04 \text{ mol}$ $n(\text{O}_2) = 0.02 \text{ mol}$ $\text{Volume of O}_2 = 0.0200 \times 24\,000 = 480 \text{ cm}^3$	B1 B1 B1 B1	
7 (a)	$0.0250 \times 23.0 = 0.575 \text{ g}$	B1	
7 (b)	$2\text{Na(s)} + 2\text{H}_2\text{O(l)} \rightarrow 2\text{NaOH(aq)} + \text{H}_2\text{(g)}$	B1	
7 (c)	$n(\text{H}_2) = 0.0125 \text{ mol}$ $\text{Volume H}_2 = 0.0125 \times 24\,000 = 300 \text{ cm}^3$	B1 B1	
7 (d) (i)	$n(\text{NaOH}) = n(\text{Na}) = 0.0250 = c \times 50/1000$ $c = 0.500 \text{ mol dm}^{-3}$	B1	
7 (d) (ii)	$c = 0.500 \times 40.0 = 20.0 \text{ g dm}^{-3}$	B1	
8 (a)	58.5 g	B1	
8 (b)	$n(\text{Cl}_2) = (2.5 \times 10^3)/24 = 1.04 \times 10^2 \text{ mol}$ $n(\text{NaOH}) = 2 \times 1.04 \times 10^2 = 2.08 \times 10^2 \text{ mol}$ $2.08 \times 10^2 = 4.00 \times V \text{ (in dm}^3\text{)}$ $\therefore V = 5.2 \times 10^7 \text{ dm}^3$	B1 B1 B1	
9 (a) (i)	$\text{C} : \text{H} = 54.55/12.0 : 9.09/1.0 : 36.36/16.0 = 4.55 : 9.09 : 2.27$ $\text{Empirical formula} = \text{C}_2\text{H}_4\text{O}$	B1 B1	
9 (a) (ii)	$pV = nRT$ $n = \frac{(103 \times 10^3) \times (72.0 \times 10^{-6})}{8.314 \times 373} =$ $0.00239 \text{ mol}$ $M = \frac{0.2103}{0.00239} = 88.0$	B1 B1 B1 B1	
9 (a) (iii)	$\text{Molecular formula} = \text{C}_2\text{H}_4\text{O} \times 88/44 = \text{C}_4\text{H}_8\text{O}_2$	B1	
10 (a) (i)	$\text{Al}^{3+}$ $\text{SO}_4^{2-}$	B1 B1	

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**PT Brinkman**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the PT Brinkman text.

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## **Table of Contents Ocr Chemistry As Answers**

1. Understanding the eBook Ocr Chemistry As Answers
  - The Rise of Digital Reading Ocr Chemistry As Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Ocr Chemistry As Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ocr Chemistry As Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr Chemistry As Answers
  - Personalized Recommendations
  - Ocr Chemistry As Answers User Reviews and Ratings
  - Ocr Chemistry As Answers and Bestseller Lists
5. Accessing Ocr Chemistry As Answers Free and Paid eBooks
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  - Ocr Chemistry As Answers Enhanced eBook Features
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  - Adjustable Fonts and Text Sizes of Ocr Chemistry As Answers
  - Highlighting and Note-Taking Ocr Chemistry As Answers
  - Interactive Elements Ocr Chemistry As Answers
8. Staying Engaged with Ocr Chemistry As Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ocr Chemistry As Answers
9. Balancing eBooks and Physical Books Ocr Chemistry As Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ocr Chemistry As Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ocr Chemistry As Answers
  - Setting Reading Goals Ocr Chemistry As Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ocr Chemistry As Answers
  - Fact-Checking eBook Content of Ocr Chemistry As Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

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