

- (b) A solution is prepared by the addition of 10.0 mL of 0.10 M CH_3COOH to 10.0 mL of 0.10 M NaCH_3COO . Calculate the pH of the solution.

- (c) The solution is a buffer. Calculate the pH of the solution.

The total volume is 20.0 mL. The concentration of CH_3COOH is 0.05 M.

In this case, we should use the Henderson-Hasselbalch equation to calculate the pH of the solution.

For CH_3COOH and CH_3COO^- , $K_a = 1.8 \times 10^{-5}$.

$$[\text{CH}_3\text{COOH}] = 0.05 \text{ M} \quad [\text{CH}_3\text{COO}^-] = 0.05 \text{ M}$$

$$K_a = \frac{[\text{CH}_3\text{COO}^-][\text{H}^+]}{[\text{CH}_3\text{COOH}]} \Rightarrow [\text{H}^+] = \frac{K_a [\text{CH}_3\text{COOH}]}{[\text{CH}_3\text{COO}^-]} = \frac{1.8 \times 10^{-5} \times 0.05}{0.05} = 1.8 \times 10^{-5} \text{ M}$$

$$\text{pH} = -\log[\text{H}^+] = -\log(1.8 \times 10^{-5}) = 4.74$$

$$\text{pH} = 4.74$$

Reaction: $\text{CH}_3\text{COOH} + \text{H}_2\text{O} \rightleftharpoons \text{CH}_3\text{COO}^- + \text{H}_3\text{O}^+$
Initial: 0.10 M, 0.10 M, 0.00 M, 0.00 M
Change: -x, -x, +x, +x
Equilibrium: 0.05 M, 0.05 M, x, x
pH = 4.74

- (b) A solution is prepared by the addition of 10.0 mL of 0.10 M CH_3COOH to 10.0 mL of 0.10 M NaCH_3COO . Calculate the pH of the solution.

The total volume is 20.0 mL. The concentration of CH_3COOH is 0.05 M.



- (c) The solution is a buffer. Calculate the pH of the solution.

For CH_3COOH and CH_3COO^- , $K_a = 1.8 \times 10^{-5}$.

- (d) Calculate the pH of the solution.

$$K_a = \frac{[\text{CH}_3\text{COO}^-][\text{H}^+]}{[\text{CH}_3\text{COOH}]} \Rightarrow [\text{H}^+] = \frac{K_a [\text{CH}_3\text{COOH}]}{[\text{CH}_3\text{COO}^-]} = \frac{1.8 \times 10^{-5} \times 0.05}{0.05} = 1.8 \times 10^{-5} \text{ M}$$

$$\text{pH} = -\log[\text{H}^+] = -\log(1.8 \times 10^{-5}) = 4.74$$

$$\text{pH} = 4.74$$

$$\text{pH} = 4.74$$

$$\text{pH} = 4.74$$

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