

Directed Reading for
Content Mastery**Key Terms**
Chemical Bonds

Directions: Match each term in Column I with its description in Column II. Write the letter of the correct term in the space provided.

Column I

- _____ 1. binary compound
- _____ 2. chemically stable
- _____ 3. nonpolar molecule
- _____ 4. ion
- _____ 5. ionic bond
- _____ 6. polar molecule
- _____ 7. oxidation number
- _____ 8. chemical formula
- _____ 9. covalent bond
- _____ 10. hydrate
- _____ 11. chemical bond
- _____ 12. polyatomic ion

Column II

- a. number that indicates how many electrons an atom must gain, lose, or share to become stable
- b. shorthand that tells what elements a compound contains and the exact number of atoms of each element in a unit of the compound
- c. positively or negatively charged, covalently bonded group of atoms
- d. compound composed of two elements
- e. describes an atom that has a full outermost energy level
- f. molecule that has a slightly positive end and a slightly negative end
- g. the attraction that forms between atoms when they share electrons
- h. the force that holds atoms together in a compound
- i. a compound that has water chemically attached to it
- j. the force of attraction between the opposite charges of the ions in an ionic compound
- k. molecule made of two identical atoms that share the electrons equally
- l. a charged particle that has either more or fewer electrons than protons

Physical Science Chapter 21 Section 21 3 Worksheet Answers

Patrick Vollmar



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Table of Contents Physical Science Chapter 21 Section 21 3 Worksheet Answers

1. Understanding the eBook Physical Science Chapter 21 Section 21 3 Worksheet Answers
 - The Rise of Digital Reading Physical Science Chapter 21 Section 21 3 Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Chapter 21 Section 21 3 Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Chapter 21 Section 21 3 Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Chapter 21 Section 21 3 Worksheet Answers
 - Personalized Recommendations
 - Physical Science Chapter 21 Section 21 3 Worksheet Answers User Reviews and Ratings
 - Physical Science Chapter 21 Section 21 3 Worksheet Answers and Bestseller Lists
5. Accessing Physical Science Chapter 21 Section 21 3 Worksheet Answers Free and Paid eBooks
 - Physical Science Chapter 21 Section 21 3 Worksheet Answers Public Domain eBooks
 - Physical Science Chapter 21 Section 21 3 Worksheet Answers eBook Subscription Services
 - Physical Science Chapter 21 Section 21 3 Worksheet Answers Budget-Friendly Options
6. Navigating Physical Science Chapter 21 Section 21 3 Worksheet Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Physical Science Chapter 21 Section 21 3 Worksheet Answers Compatibility with Devices
 - Physical Science Chapter 21 Section 21 3 Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Chapter 21 Section 21 3 Worksheet Answers
 - Highlighting and Note-Taking Physical Science Chapter 21 Section 21 3 Worksheet Answers
 - Interactive Elements Physical Science Chapter 21 Section 21 3 Worksheet Answers
 8. Staying Engaged with Physical Science Chapter 21 Section 21 3 Worksheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Chapter 21 Section 21 3 Worksheet Answers
 9. Balancing eBooks and Physical Books Physical Science Chapter 21 Section 21 3 Worksheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Chapter 21 Section 21 3 Worksheet Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Physical Science Chapter 21 Section 21 3 Worksheet Answers
 - Setting Reading Goals Physical Science Chapter 21 Section 21 3 Worksheet Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Physical Science Chapter 21 Section 21 3 Worksheet Answers
 - Fact-Checking eBook Content of Physical Science Chapter 21 Section 21 3 Worksheet Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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