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Understanding Concepts

- What is a stable electron configuration?
- What does each dot in an electron dot diagram represent?
- What process changes atoms into ions?
- What keeps the ions in their fixed positions within a crystal lattice?
- What are subscripts used for in chemical formulas?
- Explain why a melted ionic compound is a good conductor of electric current, but a solid ionic compound is a poor conductor of electric current.
- What distinguishes single, double, and triple covalent bonds?
- Explain why the covalent bonds in molecules of O_2 are always nonpolar.
- Explain why, in a covalent bond between oxygen and hydrogen, the hydrogen atom has a partial positive charge and the oxygen atom has a partial negative charge.
- What is the name of the binary compound formed from potassium and iodine?
- Write the formulas for the compounds called copper(I) chloride and copper(II) chloride.
- Name the compounds represented by the space-filling models labeled A, B, and C.



- In general, what determines the strength of metallic bonds?
- What properties of copper and tin change when these metals are mixed together to form bronze?
- What advantage of magnesium is retained in magnesium alloys? What disadvantage is reduced?

Chemical Bonds 187

Assessment



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Reviewing Content

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|-------|------|------|
| 1. b | 2. c | 3. a |
| 4. d | 5. c | 6. b |
| 7. d | 8. b | 9. c |
| 10. a | | |

Understanding Concepts

- In a stable electron configuration, the highest occupied energy level is filled with electrons.
- Each dot represents a valence electron.
- The transfer of electrons.
- Attractions between neighboring cations and anions keep the ions in fixed positions within the lattice.
- A subscript is used to show the number of atoms of an element in a molecule or the ratio of ions in a crystal lattice.
- When an ionic compound melts, ions can move away from their fixed locations in the crystal lattice.
- Two atoms share two electrons in a single bond, four in a double bond, and six in a triple bond.
- The covalent bonds in molecules of elements are always nonpolar because the atoms have the same attraction for electrons.
- The oxygen atom has a greater attraction for electrons than the hydrogen atom does.
- Potassium iodide
- CuCl and CuCl_2
- A is sulfur trioxide, B is carbon monoxide, and C is nitrogen dioxide
- In general, the more valence electrons a metal has, the stronger the metal bonds are.
- A mixture of copper and tin is harder and stronger than either metal in its pure form.
- The advantage that is retained is that magnesium is a lightweight metal. The disadvantage that is reduced is that magnesium is a soft metal.

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Physical Science Chapter 1 assessment Answers

CL Gary



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