

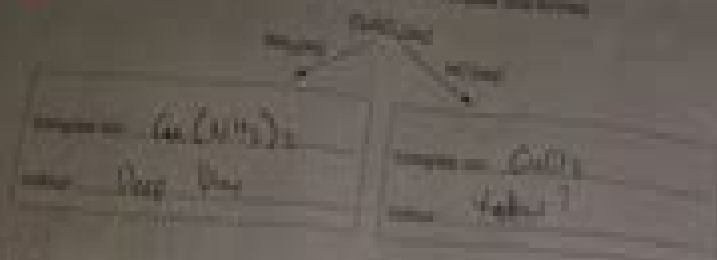
Question 1

Answer all the questions

The reaction is about the synthesis of copper complexes and ligands.

(a) The reaction shows the synthesis of a copper complex.

In the boxes, write the formula and name of the ligands and solvent.



(b) Cu^{2+} ions form a complex ion B with two ethylenediamine ligands and two water molecules. The ethylenediamine ion is a bidentate ligand.

The skeletal formula of the ethylenediamine ion is shown in Fig. 1.1 below.



Fig. 1.1

(c) What is meant by the term bidentate ligand?

A ligand that has two donor atoms, each of which can form a coordinate bond with the central metal ion.

(d) The complex ion A is shown in Fig. 1.2 below.

The structure of one of the ethylenediamine ligands is shown below. The charge has been given.



Complex A

Complete the 3D diagram of the other two ethylenediamine ligands of A. You do not need to include any charges.

Indicate with ticks whether the stereoisomers are cis, trans, optical or a combination of these types.

Structure														
Type	<table border="1"> <tr> <td>cis</td> <td>☐</td> </tr> <tr> <td>trans</td> <td>☒</td> </tr> <tr> <td>optical</td> <td>☐</td> </tr> </table>	cis	☐	trans	☒	optical	☐	<table border="1"> <tr> <td>cis</td> <td>☒</td> </tr> <tr> <td>trans</td> <td>☐</td> </tr> <tr> <td>optical</td> <td>☒</td> </tr> </table>	cis	☒	trans	☐	optical	☒
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optical	☒													

(e) What is the empirical formula, including the charge, of the complex ion A?

$[\text{Cu}(\text{NH}_3)_4(\text{eth})_2]^{2+}$

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Tobias Bleicker



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