

7. Chromite is the main ore of chromium. The chromium-containing compound in chromite is  $\text{FeCr}_2\text{O}_4$ . The percentage of chromium in a sample of chromite can be determined using the method below.

## Step 1

A 5.28 g sample of chromite ore is heated with sodium peroxide,  $\text{Na}_2\text{O}_2$ .



Water is added to the resulting mixture.

$\text{Na}_2\text{O}_2$  dissolves in the water forming a solution containing  $\text{CO}_3^{2-}$  ions.

## Step 2

The mixture from Step 1 is filtered and the filtrate is made up to 1.00 dm<sup>3</sup> in a volumetric flask.

A 25.0 cm<sup>3</sup> sample of this alkaline solution is pipetted into a conical flask and an excess of aqueous potassium iodide is added.

- A redox reaction takes place between  $\text{I}^-$  ions,  $\text{CrO}_4^{2-}$  ions and  $\text{H}_2\text{O}$ .
- In this reaction 1 mol  $\text{CrO}_4^{2-}$  reacts 1.5 mol  $\text{I}_2$ .

## Step 3

The resulting mixture is titrated with 0.100 mol dm<sup>-3</sup> sodium thiosulfate,  $\text{Na}_2\text{S}_2\text{O}_3$ , to estimate the  $\text{I}_2$  present.



The average titre of  $\text{Na}_2\text{S}_2\text{O}_3$  was 26.5 cm<sup>3</sup>.

- (a) In Step 1,  $\text{Na}_2\text{O}_2$  and  $\text{NaFeO}_2$  react with water forming an alkaline solution containing a brown precipitate. This is not a redox reaction.

Write equations for:

- the reaction of  $\text{Na}_2\text{O}_2$  with water
- the reaction of  $\text{NaFeO}_2$  with water.



$$\begin{array}{rcl} \text{I}^- + 2 + (2 \times 2) & = & 0 \\ \text{I}^- & - & 1 \\ & & 2 \end{array}$$

- (b) Calculate the percentage, by mass, of chromium in the ore.

Give your answer to one decimal place.

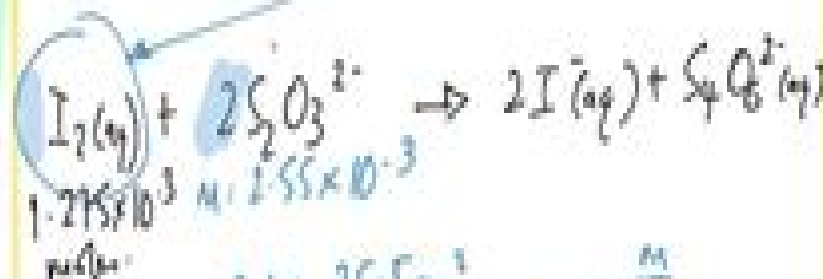
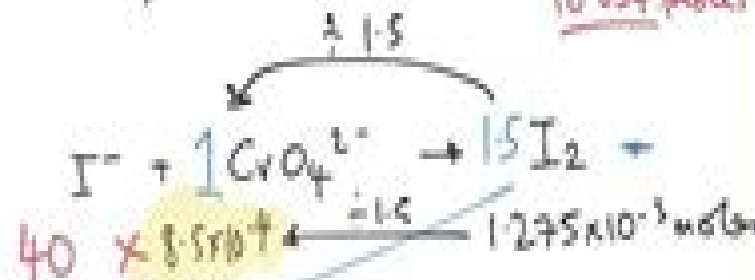
$$\% \text{ Cr} = \frac{\text{mass Cr}}{\text{mass ore}} \times 100$$

$$\text{mass} = \text{moles} \times \text{M.R.}$$

$$1.76 \quad 0.034 \times 52.0$$



0.034 moles



$$0.1 \times \frac{25.5 \text{ cm}^3}{1000}$$

$$C = \frac{n}{V}$$

$n = CV$

- (c) This part refers to Step 2 of the method.

Write the redox reaction between  $\text{I}^-$  ions,  $\text{CrO}_4^{2-}$  ions and  $\text{H}_2\text{O}$ .

- $\text{CrO}_4^{2-}$  ions are reduced to chromate(VI) ions,  $\text{Cr}^{3+}$ .
- $\text{I}^-$  ions are oxidised to make  $\text{I}_2$ .
- Construct an overall equation for the redox reaction and write half equations for the oxidation and reduction.

Overall equation:

Half equations:

# Ocr Chemistry F325 Stretch And Challenge

**Mike Smith**



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OCR(A) A2 Chemistry Student Unit Guide: Unit F325 Equilibria, Energetics and Elements Mike Smith,2009-06-26

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