

16. The dissociation of water is shown below:



- At 25 °C, the auto product of water, K_w , is 1.0×10^{-14} mol² dm⁻⁶.
At 65 °C, the auto product of water, K_w , is 1.0×10^{-13} mol² dm⁻⁶.

17. Explain what the dissociation of water is an indication of water's behavior.

Endothermic.

18. Predict, using a calculation, whether a solution of 1.0 mol dm⁻³ is acidic, basic or neutral. (3)

Acidic

Explain by calculating pH.

19. pK_a , pH and pOH are significant notions.

Calculate pK_a of H_2O .

Give your answer to two decimal places.

$$-\log(1.0 \times 10^{-14}) = 14.00$$

$$pK_a = 14.00$$

20. Describe the autoionization of water in terms of water and the resulting species in solution at 25 °C.

Calculate the pH of the resulting solution of water at 25 °C.

Give your answer to two decimal places.

$$pH = 7.00$$

pH = _____ (3)

(Total: 18)

Ocr Chemistry A2 Unofficial Mark Scheme 2014

Richard Parsons



Ocr Chemistry A2 Unofficial Mark Scheme 2014:

This book delves into Ocr Chemistry A2 Unofficial Mark Scheme 2014. Ocr Chemistry A2 Unofficial Mark Scheme 2014 is a vital topic that needs to be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Ocr Chemistry A2 Unofficial Mark Scheme 2014, encompassing both the fundamentals and more intricate discussions.

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 - Chapter 3: Ocr Chemistry A2 Unofficial Mark Scheme 2014 in Everyday Life
 - Chapter 4: Ocr Chemistry A2 Unofficial Mark Scheme 2014 in Specific Contexts
 - Chapter 5: Conclusion
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