

Configuration du serveur du système Polycom® CMA™

Ce chapitre décrit les différentes opérations à effectuer pour configurer le serveur d'un système Polycom® Converged Management Application™ (CMA™) et le connecter au réseau.



Si vous installez une configuration redondante de Polycom CMA, exécutez toutes les procédures figurant dans ce chapitre sur les deux serveurs, sauf instruction contraire.

Réunir les éléments nécessaires

Avant d'installer un système Polycom CMA, vous devez réunir les différents éléments suivants :

- *Notes de diffusion du système Polycom CMA*
- Serveur du système Polycom CMA
- Étude complète du site ou boîte à outils du projet
- Ordinateur comportant un port série ou un port Ethernet

Si vous souhaitez vous connecter au serveur du système Polycom CMA par un port Ethernet, vous devez disposer également des composants suivants :

- Un câble croisé OU
- Un concentrateur ou un commutateur et deux câbles Ethernet

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R Sanford



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web donor whereas a base generically represented by b is a proton acceptor when acids and bases combine they swap a proton $h_a + b \rightarrow h_b$ strong acids e.g. sulfuric acid have a very pronounced tendency to transfer a proton to a base whereas weak acids

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web answer based on the data the strong acid strong base combination of hydrochloric acid and sodium hydroxide had the neutral pH at the equivalence point being 7 this result is expected because a combination of strong acids and strong bases should result in a neutral at the equivalence point

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web 1 approximately 3 g of hydrated oxalic acid $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ was weighted to the nearest 0.01 g in a 50 ml beaker 2 about 30 ml of distilled water was added to the beaker to dissolve the oxalic acid 3 the solution was transferred into 250 ml volumetric flask the beaker was rinsed and its contents were poured into the flask

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web solution a titrant volume 0 ml the solution ph is due to the acid ionization of HCl because this is a strong acid the ionization is complete and the hydronium ion molarity is 0.100 M the ph of the solution is then $\text{pH} = -\log 0.100 = 1.000$ b titrant volume 12.50 ml

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