

Periodic Trends Worksheet

Directions: Use the trends in electronegativity to answer the questions.

1. Write the following elements by increasing atomic radius: sodium, sulfur, chlorine, oxygen, potassium.

K, Na, S, Cl, O

2. Write the following elements by increasing electronegativity: sodium, oxygen, sulfur, chlorine, potassium.

O, Cl, S, Na, K

3. Which atom has a higher first ionization energy than sodium?

Chlorine (because it has more protons than sodium and is closer to the nucleus)

4. Which two elements in the same family generally have similar properties?

Sulfur and selenium (because they are both in the same family and have similar properties)

5. Arrange the following properties in order from left to right across the periodic table:

- a. atomic radius (increases to the right) **increases**
- b. first ionization energy **increases**
- c. electronegativity **increases**

6. Which period in the periodic table shows atoms with the greatest atomic radius? Which element has the greatest?

Period 7, francium (because it has the most protons and is furthest from the nucleus)

7. Which period in the periodic table shows atoms with the smallest atomic radius? Which element has the smallest?

Period 1, hydrogen (because it has the fewest protons and is closest to the nucleus)

8. Write the atoms in order from the lowest to the highest atomic radius:

a. **Br**
 b. **As**
 c. **Cl**
 d. **Se**
 e. **I**
 f. **Sb**

a. **Br**
 b. **As**
 c. **Cl**
 d. **Se**
 e. **I**
 f. **Sb**

9. Write the atoms in order from the lowest to the highest electronegativity:

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Richard Bailey

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the author's name.

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