

**Physical Science – 2nd Semester – Final Exam Study Guide – (2012-2013)****1. Physical Science Basics**

- a. What tool(s) would you want to use to find the:
- i. Mass of an object? triple beam balance Basic SI Unit? kg
  - ii. Volume of an object? ruler or water displacement method Basic SI Unit? cm<sup>3</sup> or ml
  - iii. Length of an object? Ruler Basic SI Unit? meter
  - iv. Density of an object? graduated cylinder Basic SI Unit? ml
- b. What is the formula for calculating density? mass/volume
- i. What is the density of a cube that is 15 g and has a volume of 3 cm<sup>3</sup>? 5g/cm<sup>3</sup>
  - ii. Compare the density of solid H<sub>2</sub>O to liquid H<sub>2</sub>O, and explain what observable effect this has.  
solid H<sub>2</sub>O is less dense than liquid H<sub>2</sub>O

**2. Chemistry Basics**

- a. Element (def.) substance made up of one kind of atom
- b. Atom (def.) basic building block of matter
- c. Compound (def.) a molecule that contains atoms or molecules of two or more different elements
- d. Molecule (def.) a combination of two or more of the same or different atoms that are bonded together

**3. Model of the Atom: Evolution of the Atomic Theory**

Atoms are so small that many scientists throughout history have created models to describe them.

Why have atomic theories changed over time? Better technology came around.

Scientist	Year	Major Discovery/ Contribution	Name of atomic model
John Dalton	<u>1802</u>	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.	<u>Billiard Ball Model</u>
J.J. Thomson	<u>1897</u>	Negatively charged electron	<u>Plum Pudding Model</u>
Niels Bohr	<u>1913</u>	Electrons move <u>fixed orbits</u> around the atom's nucleus in specific layers, or <u>electron shells</u> ; Atoms absorb or give off energy when the electrons move from one shell ( <u>energy level</u> ) to another	<u>Bohr Model</u>
Ernest Rutherford	<u>1909</u>	1909 ( <u>Based on idea that positive reflects positive</u> ) atom mostly empty space; electrons orbit randomly around positively charged nucleus	<u>Planetary Model</u>

**4. Periodic Table**

- a. The current periodic table is arranged by increasing atomic number.
- b. **Reading the Periodic Table**
- i. An element's properties can be predicted from its location in the periodic table.
  - ii. The atomic number represents the number of protons.
  - iii. The atomic mass represents the number of neutrons and protons.

# Physical Science 2nd Semester Study Guide 2012 2013

**Katrin Zwirgmaier**



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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) + \text{C}_2\text{H}_2(\text{g}) \rightarrow \text{O}_2(\text{g}) + \text{C}_2\text{H}_2(\text{g}) + 2 \text{ H}_2(\text{g}) + \text{C}_2\text{H}_6(\text{g})$   $\Delta h = 94.5 \text{ kJ}$   
 $\text{H}_2\text{O}(\text{g}) + \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{g})$   $\Delta h = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g})$   $\Delta h = 283 \text{ kJ}$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l})$   $\Delta h = 1411 \text{ kJ/mole}$   
 $\text{C}_2\text{H}_6(\text{g}) + 7 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l})$   $\Delta h = 1560 \text{ kJ/mole}$   
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$   $\Delta h = 285 \text{ kJ/mole}$

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web thus for reaction 9 2 in which 1 mole of  $\text{H}_2\text{O}$  and 1 mole of  $\text{OH}^-$  combine to form 2 moles of  $\text{H}_2\text{O}$   $\Delta H = -55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$  is positive if heat is absorbed that is if  $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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