

QUESTION 3

3. 1. Learners made a mini volcano in science laboratory by adding sodium bicarbonate to ethanoic acid. They added 100 ml of a $0,2 \text{ mol}\cdot\text{dm}^{-3}$ ethanoic acid solution to 10 g of NaHCO_3 to start the reaction of the volcano.

The balanced equation for this reaction is:



- 3.1.1 Define the term limiting reagent. (2)
3.1.2 Determine the limiting reagent in this reaction. (6)
3.1.3 Calculate the mass of the other substance in excess. (3)
3.1.4 Calculate the volume of CO_2 produced at STP. (4)

3.2 A solution of potassium hydroxide (KOH) is prepared by dissolving 3,36 g crystals of KOH in 250 cm^3 of water.

Calculate the concentration of the potassium hydroxide solution. (4)

3.3 25 cm^3 of a potassium hydroxide solution of concentration $0,25 \text{ mol}\cdot\text{dm}^{-3}$ completely neutralises a dilute solution of sulphuric acid (H_2SO_4) in a flask.

The unbalanced equation below represents the reaction that takes place:



- 3.3.1 Balance the above equation. (2)
3.3.2 Calculate the mass of sulphuric acid in the flask. (5)

[25]

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