

OCR A LEVEL CHEMISTRY DEFINITIONS

Module 2:	
Term	Definition
Isotope	atoms of the same element with different numbers of neutrons and different masses
Relative isotopic mass	Mass compared with 1/12th mass of carbon-12
Relative atomic mass	weighted mean mass compared with 1/12th mass of carbon-12 atom.
Empirical formula	the simplest whole number ratio of atoms of each element present in a compound
Molecular formula	the number and type of atoms of each element in a molecule
Acid/Bronsted Lowry Acid	release H ⁺ ions in aqueous solution/Proton donor
Alkalis/ Bronsted Lowry Base	release OH ⁻ ions in aqueous solution/proton acceptor
Strong acid	Fully dissociates to release H ⁺ ions in aqueous solution
Weak acid	Partially dissociates to release H ⁺ ions in aqueous solution
Orbital	a region around the nucleus that can hold up to two electrons, with opposite spins
Ionic Bond	electrostatic attraction between positive and negative ions
Covalent Bond	strong electrostatic attraction between a shared pair of electrons and the nuclei of the bonded atoms
Metallic Bond	strong electrostatic attraction between cations (positive ions) and delocalised electrons
Electronegativity	ability of an atom to attract the bonding electrons in a covalent bond
Module 3:	
First Ionisation Energy	The energy required to remove 1 mol of electrons from 1 mol of gaseous atoms to form 1 mol of gaseous 1+ ions
Second Ionisation Energy	The energy required to remove 1 mol of electrons from 1 mol of gaseous 1+ ions to form 1 mol of gaseous 2+ ions
Disproportionation	oxidation and reduction of the same element
Standard conditions	298K, 1atm/100kPa, 1 moldm ⁻³ solutions
Standard enthalpy of formation	the enthalpy change when 1 mole of a compound is formed from its elements in their standard states under standard conditions
Standard enthalpy of combustion	the enthalpy change when 1 mole of a substance is completely burned in oxygen under standard conditions
Standard enthalpy of neutralisation	the enthalpy change when 1 mole of an acid reacts with 1 mole of an alkali to form 1 mole of water under standard conditions
Standard enthalpy of solution	the enthalpy change when 1 mole of an ionic solid is dissolved in a large volume of water under standard conditions
Standard enthalpy of hydration	the enthalpy change when 1 mole of gaseous ions is dissolved in a large volume of water under standard conditions
Standard enthalpy of atomisation	the enthalpy change when 1 mole of gaseous atoms is formed from the element in its standard state under standard conditions
Standard enthalpy of formation of an ionic compound	the enthalpy change when 1 mole of an ionic compound is formed from its elements in their standard states under standard conditions
Standard enthalpy of combustion of an ionic compound	the enthalpy change when 1 mole of an ionic compound is completely burned in oxygen under standard conditions
Standard enthalpy of neutralisation of an ionic compound	the enthalpy change when 1 mole of an ionic compound reacts with 1 mole of an alkali to form 1 mole of water under standard conditions
Standard enthalpy of solution of an ionic compound	the enthalpy change when 1 mole of an ionic compound is dissolved in a large volume of water under standard conditions
Standard enthalpy of hydration of an ionic compound	the enthalpy change when 1 mole of gaseous ions is dissolved in a large volume of water under standard conditions
Standard enthalpy of atomisation of an ionic compound	the enthalpy change when 1 mole of gaseous atoms is formed from the element in its standard state under standard conditions

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Michael Seilmaier



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