

Part 2

GCSE Chemistry OCR A (Combined Science - **HIGHER**) Paper 2 Exam Walkthrough

Combined Science Higher - Chemistry

Walkthrough Mark Paper

12
19

Look at the equation for the equilibrium reaction between sulfur dioxide, SO_2 , oxygen and sulfur trioxide, SO_3 , in a closed system.



- (a) Predict the effect of adding more oxygen to the equilibrium mixture.

Explain your answer.

(2)

- (b) Predict the effect of increasing the total pressure in the equilibrium mixture.

Explain your answer.

(2)

- (c) Predict the effect of increasing the **temperature** of the equilibrium mixture.

Explain your answer.

(2)

- (d) Sulfur dioxide for this reaction is made by burning sulfur.



Calculate the mass of sulfur needed to make 48 tonnes of sulfur dioxide.

Give your answer to 2 significant figures.

The relative atomic mass, A_r , of S is 32.1 and of O is 16.0.

Ocr Chemistry B 2013 B74

IM Harris



Ocr Chemistry B 2013 B74:

Whispering the Strategies of Language: An Emotional Journey through **Ocr Chemistry B 2013 B74**

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