

## Chapter 6: Chemical Bonding Guided Notes

### I. Why do Atoms Combine?

#### A. Chemical Formula

- The chemical formula shows:
  - The \_\_\_\_\_ in the compound.
  - The ratio of their \_\_\_\_\_



#### B. Chemical Bond

- Strong attractive force between \_\_\_\_\_ or ions in a molecule or compound.
- Formed by:
  - \_\_\_\_\_ electrons, (e-) by losing or gaining electrons.
  - \_\_\_\_\_ electrons (e-)

#### C. Stability

- Octet Rule
  - Most atoms form bonds in order to have \_\_\_\_\_ valence electrons (e-).
  - \_\_\_\_\_ outer energy level.
  - Want to be like the Noble Gases (group 18).
- Stability is the driving force behind bond formation. All atoms bond to become stable.
- Transferring e-



- Sharing e-



### II. Kinds of Chemical Bonds

#### A. Ionic Bond

- Attraction between 2 \_\_\_\_\_ charged ions.
- Ions are \_\_\_\_\_ atoms.
- A \_\_\_\_\_ is a \_\_\_\_\_ charged ion. Forms when an atom loses an electron.  
Formed by \_\_\_\_\_.
- A \_\_\_\_\_ is a \_\_\_\_\_ charged ion. Forms when an atom gains an electron.  
Formed by \_\_\_\_\_.
- Formed by \_\_\_\_\_ electrons (e-) from a \_\_\_\_\_ to a \_\_\_\_\_.
- Ions form a 3-D crystal lattice.



# Physical Science Chapter 6 Chemical Bonding Wordwise Answers

**CH Cherryholmes**



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