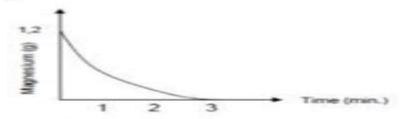
Grade 11 Physical Science M...

records the change in the mass of magnesium as the reaction proceeds at regular intervals. The balanced chemical equation for the reaction is:

$$Mg(s) + 2HCl(aq) \rightarrow MgCl_2(aq) + H_2(g)$$

The change in the mass of magnesium during the reaction is shown on the graph below.



- 6.2.1 Identify the limiting agent in this reaction. Give a reason for the answer.
- 622 Calculate the number of moles of unreacted hydrochloric acid in the test tube after 3 minutes.

Hide Answers

QUESTION/VRAAG 6

- 6.1.1 2H₂O₂ (aq) ✓ → 2H₂O (ℓ) + O₂ (q) ✓
- 6.1.2 The catalyst lowers the activation energy of the reaction < Accept: catalyst speeds up the reaction 'n Katalisator verlaag die aktiveringsenergie van die reaksie 🗸 🗸
 - Aanvaar: katalisator laat die reaksie vinniger plaasvind
- 6.1.3 OPTION 1/OPSIE 1

$$n = \frac{m}{M} \checkmark$$

$$0.049 = \frac{m}{34}$$

OPTION 2/OPSIE 2

From the balanced equation: Vanaf gebalanseerde vergelyking.

68g H₂O₂ → 24,45 dm³ O₂ ✓ ✓ X a H₂O₂ → 600 x 10⁻³ dm³ ✓ X = 00 x 24,45 68×0,6 ~

(2)

(7)[19]

(2)

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