

# F321 Revision of definitions to be learnt for exam

|                                                  |                                                                                                                                  |
|--------------------------------------------------|----------------------------------------------------------------------------------------------------------------------------------|
| <b>Isotopes</b>                                  | Atoms of the same element with different numbers of neutrons and different masses.                                               |
| <b>Relative atomic mass, <math>A_r</math></b>    | The weighted mean mass of an atom of an element compared with one-twelfth of the mass of an atom of $^{12}\text{C}$ .            |
| <b>Relative isotopic mass</b>                    | The mass of an atom of an isotope compared with one-twelfth of the mass of an atom of carbon-12.                                 |
| <b>Relative formula mass</b>                     | The relative formula mass of a compound is equal to the sum of the individual relative atomic masses.                            |
| <b>Relative molecular mass, <math>M_r</math></b> | The weighted mean mass of a molecule compared with one-twelfth of the mass of an atom of carbon-12.                              |
| <b>Mole</b>                                      | The amount of any substance containing as many particles as there are carbon atoms in exactly 12 g of the carbon-12 isotope.     |
| <b>Molecular formula</b>                         | The number of atoms of each element in a molecule.                                                                               |
| <b>Empirical formula</b>                         | The simplest whole-number ratio of atoms of each element present in a compound.                                                  |
| <b>Acid</b>                                      | A species that is a proton donor.                                                                                                |
| <b>Base</b>                                      | A species that is a proton acceptor.                                                                                             |
| <b>Concentrated Solution</b>                     | Solutions containing a higher number of moles of solute in a given volume of solution                                            |
| <b>Dilute Solution</b>                           | Solutions containing a fewer number of moles of solute in a given volume of solution                                             |
| <b>Hydrated</b>                                  | Salt containing water molecules known as "water of crystallisation"                                                              |
| <b>Anhydrous</b>                                 | A substance that contains no water of crystallisation                                                                            |
| <b>Oxidation</b>                                 | Loss of electrons, or an increase in oxidation number.                                                                           |
| <b>Reduction</b>                                 | Gain of electrons, or a decrease in oxidation number.                                                                            |
| <b>Redox reaction</b>                            | A reaction in which both reduction and oxidation take place.                                                                     |
| <b>(First) ionisation energy</b>                 | The energy required to remove one electron from each atom in one mole of gaseous atoms to form one mole of gaseous $1+$ ions.    |
| <b>(Second) ionisation energy</b>                | The energy required to remove one electron from each ion in one mole of gaseous $1+$ ions to form one mole of gaseous $2+$ ions. |

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**M Lipman**



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