

| Question number | Answer | Marks | Guidance |
|-----------------|---|----------------------|----------|
| 6 (c) | $M(\text{NaNO}_3) = 85.0 \text{ g mol}^{-1}$ $n(\text{NaNO}_3) = 0.04 \text{ mol}$ $n(\text{O}_2) = 0.02 \text{ mol}$ $\text{Volume of O}_2 = 0.0200 \times 24\,000 = 480 \text{ cm}^3$ | B1 B1 B1 B1 | |
| 7 (a) | $0.0250 \times 23.0 = 0.575 \text{ g}$ | B1 | |
| 7 (b) | $2\text{Na(s)} + 2\text{H}_2\text{O(l)} \rightarrow 2\text{NaOH(aq)} + \text{H}_2\text{(g)}$ | B1 | |
| 7 (c) | $n(\text{H}_2) = 0.0125 \text{ mol}$ $\text{Volume H}_2 = 0.0125 \times 24\,000 = 300 \text{ cm}^3$ | B1 B1 | |
| 7 (d) (i) | $n(\text{NaOH}) = n(\text{Na}) = 0.0250 = c \times 50/1000$ $c = 0.500 \text{ mol dm}^{-3}$ | B1 | |
| 7 (d) (ii) | $c = 0.500 \times 40.0 = 20.0 \text{ g dm}^{-3}$ | B1 | |
| 8 (a) | 58.5 g | B1 | |
| 8 (b) | $n(\text{Cl}_2) = (2.5 \times 10^3)/24 = 1.04 \times 10^2 \text{ mol}$ $n(\text{NaOH}) = 2 \times 1.04 \times 10^2 = 2.08 \times 10^2 \text{ mol}$ $2.08 \times 10^2 = 4.00 \times V \text{ (in dm}^3\text{)}$ $\therefore V = 5.2 \times 10^7 \text{ dm}^3$ | B1 B1 B1 | |
| 9 (a) (i) | $\text{C} : \text{H} = 54.55/12.0 : 9.09/1.0 : 36.36/16.0 = 4.55 : 9.09 : 2.27$ $\text{Empirical formula} = \text{C}_2\text{H}_4\text{O}$ | B1 B1 | |
| 9 (a) (ii) | $pV = nRT$ $n = \frac{(103 \times 10^3) \times (72.0 \times 10^{-6})}{8.314 \times 373} =$ $0.002\,39 \text{ mol}$ $M = \frac{0.2103}{0.00239} = 88.0$ | B1 B1 B1 B1 | |
| 9 (a) (iii) | $\text{Molecular formula} = \text{C}_2\text{H}_4\text{O} \times 88/44 = \text{C}_4\text{H}_8\text{O}_2$ | B1 | |
| 10 (a) (i) | Al^{3+} SO_4^{2-} | B1 B1 | |

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