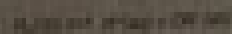


16. The dissociation of water is shown below:



At 25 °C, the ion product of water, K_w , is 1.0×10^{-14} .

At 25 °C, the ion product of water, K_w , is 1.0×10^{-14} .

17. Explain what the dissociation of water is an indication of water's behavior.

Endothermic

18. Predict, using a calculation, whether a solution of 1.0 M HCl is acidic, basic, or neutral. (1)

Acidic

Explain by calculating pH.

19. pK_a , pH , and pOH are significant values.

Calculate pK_a of HCl .

Give your answer to two decimal places.

$$-\log(4.31 \times 10^{-14}) = 13.07$$

$$pK_a = 13.03 \quad (1)$$

20. A solution of hydrochloric acid is shown with water and the resulting ions in solution at 25 °C.

Calculate the pH of the resulting solution at 25 °C.

Give your answer to two decimal places.

$$pH = 1.30$$

pH = _____ (1)

(Total: 10)

Ocr Chemistry A2 Unofficial Mark Scheme 2014

Ian Pickup



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